

CHAPTER 4

STUDY GUIDE

The Structure of the Atom

Section 4.1 Early Ideas About Matter: If the statement is true, write *true*. If the statement is false, change the *italicized* word and make it true.

- _____ 1. Ancient philosophers *regularly* performed controlled experiments.
- _____ 2. *Rutherford* suggested that matter is made up of atoms.
- _____ 3. *Democritus's* atomic theory stated that atoms separate, combine, or rearrange in chemical reactions.
- _____ 4. *Dalton's* atomic theory stated that matter is mostly empty space.
- _____ 5. Dalton was *incorrect* in thinking that atoms could not be divided into smaller particles.
- _____ 6. Dalton's atomic theory stated that atoms of different elements combine in simple whole-number ratios to form *mixtures*.
- _____ 7. Dalton thought that all atoms of a specific element have the same number of *electrons*.

Section 4.2 Defining the Atom: For each item in Column A, write the letter of the matching item in Column B.

8		Determined the mass-to-charge ratio of an electron	A. Rutherford
9		Proposed the nuclear atomic model	B. Millikan
10		Calculated the mass of an electron	C. Thomson

Complete the following table of proton, electron, and neutron characteristics.

	Particle	Symbol	Location	Relative Charge	Relative Mass
11	Proton				
12	Electron				
13	Neutron				

Section 4.3 How Atoms Differ: If the statement is true, write *true*. If the statement is false, change the *italicized* word and make it true.

- _____ 14. The number of *electrons* in an atom is referred to as its atomic number.
- _____ 15. The periodic table is arranged by *decreasing* atomic number.
- _____ 16. Atomic number is equal to the number of *electrons* in an atom.
- _____ 17. The number of *neutrons* in an atom identifies it as an atom of a particular element.
- _____ 18. *Atoms* have either a positive or a negative charge.

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RJ Shavelson



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