
AQA A Level Chemistry

- Paper 3

what colour is $[\text{Fe}(\text{H}_2\text{O})_6]^{2+}$ - ANSWER pale green

Transition Metal definition - ANSWER D block elements that form at least one stable ion with a partially filled d orbital

Classic elements which are not transition metals despite being D-Block elements - ANSWER Scandium and Zinc

properties of transition metals which are NOT shared by other metals - ANSWER -
multiple oxidation states
-coloured ions in solution
-great catalysts
-form complexes with ligands

ligand - ANSWER a molecule which forms a dative bond with a transition metal

complex ion - ANSWER A transition metal ion bonded to one or more ligands by coordinate bonds

co-ordination number - ANSWER The number of co-ordinate bonds formed to a central metal ion

What is an example of incomplete ligand substitution? - ANSWER $[\text{Cu}(\text{NH}_3)_6]^{2+}$ to $[\text{Cu}(\text{NH}_3)_4(\text{H}_2\text{O})_2]^{2+}$

why does the coordination number not change when substituting water for ammonia? - ANSWER because they are of a similar size

Why does incomplete ligand substitution usually occur? - ANSWER when substituting ligands of a different size

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