

Unit 7 Gas Laws Test Review - Answer Key

1. A 0.4 L sample of gas is collected at a temperature of 45°C. At what temperature will the gas fill a 0.725 L container?

LIST GIVENS	FORMULA	GVENS PLUGGED INTO FORMULA:
$P_1 =$ $V_1 = 0.4 \text{ L}$ $T_1 = 318 \text{ K}$ $T =$ $V_2 = 0.725 \text{ L}$ $T_2 = ?$	$\frac{V_1}{T_1} = \frac{V_2}{T_2}$	$\frac{0.4}{318} = \frac{.725}{T_2}$ $(.725 \times 318) / 0.4 = T_2$
answer with correct units: $T_2 = 576.38 \text{ K}$		

2. If 3 moles of a gas at a pressure of 3.5 atm has a volume of 110 L, what is the temperature?

LIST GIVENS	FORMULA	GVENS PLUGGED INTO FORMULA:
$P = 3.5 \text{ atm}$ $V = 110 \text{ L}$ $n = 3 \text{ mol}$ $R = 0.0821$ $T = ?$	$PV = nRT$	$(3.5 \times 110) = (3 \times 0.0821 \times T)$ $385 / .2463 = T$
answer with correct units: $T = 1563.13 \text{ K}$		

3. Identify if the following would cause an INCREASE or DECREASE in pressure in a gaseous system.

- Increase in the size of the container = **pressure decreases**
- More of the same gas is added to the container = **pressure increases**
- Increase in temperature = **pressure increases**
- A different gas is added to the container = **pressure increases**

4. A sealed container of Oxygen gas has a volume of 356 L. If heated from a temperature of 293-K to 318-K, what will the pressure be if the original pressure was 1.2 atm?

LIST GIVENS	FORMULA	GVENS PLUGGED INTO FORMULA:
$P_1 = 1.2 \text{ atm}$ $V =$ $T_1 = 293 \text{ K}$ $P_2 = ?$ $T =$ $T_2 = 318 \text{ K}$	$\frac{P_1}{T_1} = \frac{P_2}{T_2}$	$\frac{1.2}{293} = \frac{P_2}{318}$ $(1.2 \times 318) / 293 = P_2$

Gas Laws Test Review Answers

Victor M. Corman



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