

Boyle's Law Problems:

$$P_1V_1 = P_2V_2$$

$$1 \text{ atm} = 760.0 \text{ mm Hg} = 101.3 \text{ kPa} = 760.0 \text{ torr}$$

1. If 22.5 L of nitrogen at 748 mm Hg are compressed to 725 mm Hg at constant temperature. What is the new volume?
2. A gas with a volume of 4.0L at a pressure of 205kPa is allowed to expand to a volume of 12.0L. What is the pressure in the container if the temperature remains constant?

Charles' Law Problems:

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

$$1 \text{ atm} = 760.0 \text{ mm Hg} = 101.3 \text{ kPa} = 760.0 \text{ torr}$$

3. Calculate the decrease in temperature when 6.00 L at 20.0 °C is compressed to 4.00 L.
4. A container containing 5.00 L of a gas is collected at 100 K and then allowed to expand to 20.0 L. What must the new temperature be in order to maintain the same pressure (as required by Charles' Law)?

Combined Gas Law Problems:

$$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$$

$$1 \text{ atm} = 760.0 \text{ mm Hg} = 101.3 \text{ kPa} = 760.0 \text{ torr}$$

5. A gas balloon has a volume of 106.0 liters when the temperature is 45.0 °C and the pressure is 740.0 mm of mercury. What will its volume be at 20.0 °C and 780.0 mm of mercury pressure?
6. If 10.0 liters of oxygen at STP are heated to 512 °C, what will be the new volume of gas if the pressure is also increased to 1520.0 mm of mercury?

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and liquids The equation is based on a modification of the ideal gas law and approximates the behavior of real fluids taking into account the nonzero size of molecules and the attraction between them The van der Waals isotherms the model correctly predicts a mostly incompressible liquid phase but the oscillations in the phase transition zone do not fit experimental data The equation uses the following state variables the pressure of the fluid p total volume of the container containing the fluid V number of moles n and absolute temperature of the system T One form of the equation is where v is the volume of the container shared between each particle n is the total number of particles and k_B is Boltzmann's constant given by the universal gas constant R and Avogadro's constant N_A Extra parameters are introduced a is a measure for the attraction between the particles and b is the average volume excluded from v by a particle The equation can be cast into the better known form where a is a measure of the attraction between the particles b is the volume excluded by a mole of particles A careful distinction

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