

Gas Law Stoichiometry Worksheet

Directions: Use significant figures and units in the problems below.

$$PV = nRT$$

SHOW ALL WORK!!!

1. Given the following unbalanced chemical equation for the combination reaction of sodium metal and chlorine gas:



- a. What volume of chlorine gas, measured at STP, is necessary for the complete reaction of 4.81 g of sodium metal.

$$4.81 \text{ g Na} \left| \frac{1 \text{ mol}}{23.0 \text{ g Na}} \right| \left| \frac{1 \text{ mol Cl}_2}{2 \text{ mol Na}} \right| \left| \frac{22.4 \text{ L}}{1 \text{ mol}} \right| = 2.34 \text{ L Cl}_2$$

$$2.34 \text{ L Cl}_2$$



- a. Balance the above reaction.
 b. What volume of oxygen gas at 25°C and 1.04 atm is needed for the complete combustion of 5.53 g of propane?

$$n = x$$

$$V = O_2(x)$$

$$T_1 = 25 + 273 = 298 \text{ K}$$

$$P_1 = 1.05 \text{ atm}$$

$$V = \left(\frac{5.53 \text{ g C}_3\text{H}_8}{44 \text{ g}} \right) \left(\frac{1 \text{ mol C}_3\text{H}_8}{1 \text{ mol C}_3\text{H}_8} \right) \left(\frac{5 \text{ mol O}_2}{2 \text{ mol C}_3\text{H}_8} \right) = 0.628 \text{ moles}$$

$$V = 14.6 \text{ L}$$

$$14.6 \text{ L}$$

3. Potassium permanganate, KMnO_4 , is produced commercially by oxidizing aqueous potassium manganate, K_2MnO_4 .



- a. Balance the above reaction.
 b. What volume of $\text{Cl}_2(g)$, measured at STP, is needed to produce 10.0 g of KMnO_4 ?

$$V = \frac{nRT}{P}$$

$$10.0 \text{ g KMnO}_4 \left| \frac{1 \text{ mol}}{158 \text{ g}} \right| \left| \frac{1 \text{ mol Cl}_2}{2 \text{ mol}} \right| \left| \frac{22.4 \text{ L}}{1 \text{ mol Cl}_2} \right| = 0.709 \text{ L Cl}_2$$

$$0.709 \text{ L}$$

4. If water is added to magnesium nitride, ammonia gas is produced when the mixture is heated.



- a. Balance the above reaction.
 b. If 10.3 g of magnesium nitride is treated with 10.3 g of water, what volume of ammonia gas would be collected at 24°C and 752 mmHg?

$$10.3 \text{ g Mg}_3\text{N}_2 \left| \frac{1 \text{ mol}}{100.9 \text{ g}} \right| \left| \frac{2 \text{ mol NH}_3}{1 \text{ mol}} \right| = 0.204 \text{ mol}$$

Lim Reagent

$$10.3 \text{ g H}_2\text{O} \left| \frac{1 \text{ mol}}{18 \text{ g}} \right| \left| \frac{2 \text{ mol NH}_3}{3 \text{ mol}} \right| = 0.381 \text{ mol}$$

$$24 + 273 = 297 \text{ K}$$

$$752 / 760 = 0.989 \text{ atm}$$

$$V = \frac{nRT}{P} = \frac{(0.204 \text{ mol} + 0.082 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K}) (297 \text{ K})}{0.989 \text{ atm}} = 5.03 \text{ L}$$

$$5.03 \text{ L}$$

Gas Stoichiometry Practice Sheet Answer Key

Daniel F McAuley



Gas Stoichiometry Practice Sheet Answer Key:

Unveiling the Power of Verbal Artistry: An Psychological Sojourn through **Gas Stoichiometry Practice Sheet Answer Key**

In a global inundated with monitors and the cacophony of immediate communication, the profound energy and psychological resonance of verbal art usually disappear in to obscurity, eclipsed by the constant onslaught of noise and distractions. However, located within the lyrical pages of **Gas Stoichiometry Practice Sheet Answer Key**, a captivating perform of literary elegance that pulses with organic feelings, lies an memorable journey waiting to be embarked upon. Written by way of a virtuoso wordsmith, that enchanting opus instructions viewers on a mental odyssey, gently revealing the latent possible and profound impact stuck within the elaborate web of language. Within the heart-wrenching expanse with this evocative evaluation, we can embark upon an introspective exploration of the book is main themes, dissect their interesting writing model, and immerse ourselves in the indelible impact it leaves upon the depths of readers souls.

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