

## Chapter 5: Gases - Things to know\*

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*Please remember that simply knowing or memorizing the information below is not sufficient. You must be able to know when and how to APPLY the items below appropriately.*

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### **5.1 Basics**

Pressure and Units *(see guide!)*

### **5.2 Simple Gas Laws** *(see guide!)*

### **5.3 Ideal Gas Law** *(see preview below!)*

Used to determine one gas variable for a given scenario, knowing the other variables.

A.  $PV=nRT$  ( $R = 0.0821 \text{ L atm/mol K}$ ;  $8.314 \text{ L kPa/mol K}$ ; be sure to match all units to the constant)

B. Limitations:

1. Works well at low pressures *as the molecules of the gas rarely collide or come into close enough proximity for intermolecular forces to be significant*
2. Works well at high temperatures *as the molecules of the gas have so much energy that the intermolecular forces and energy lost in collisions is negligible*
3. Most gases do not behave ideally above 1 atm pressure
4. Does not work well near the condensation conditions of a gas
5. Deviations occur due to stronger attractive forces

C. Derivation to relate molar mass and density

$D = m/V$   $n = m/M$  where  $M$  is molar mass

$PV = m/MRT$ ;  $M = DRT/P$

### **5.4 Gas Mixtures and Collecting Gas of Water**

### **5.5 Diffusion and Effusion**

# Gases Study Guide

**Louis E. Noble, Joseph A. Roenigk**

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