

Q.57	<p>Two different enzyme catalysis reactions I and II have identical Y-intercepts for the Lineweaver-Burke (equation given below) plots. The slope for reaction I is twice than that of reaction II.</p> <p>If the initial concentrations of enzymes in I and II are same, the correct statement(s) is(are)</p>
	$\frac{1}{v} = \frac{1}{v_{max}} + \frac{K_M}{v_{max}} \frac{1}{[S]}$ <p>where <math>v</math> and <math>v_{max}</math> are rate and maximum rate; <math>K_M</math> is Michaelis-Menten constant, and <math>[S]</math> is substrate concentration.</p>
(A)	Reactions I and II have same turn over number
(B)	Michaelis-Menten constants for reactions I and II are identical
(C)	Michaelis-Menten constant for reaction I is twice than that of reaction II
(D)	The rates of the elementary steps for reactions I and II are identical

Q.58	<p>The enthalpy change for the exothermic reaction between <math>\text{BeI}_2</math> and <math>\text{HgF}_2</math> is _____ <math>\text{kJ mol}^{-1}</math> (rounded off to the nearest integer)</p> <p>(Given: Bond dissociation energy (in <math>\text{kJ mol}^{-1}</math>) for <math>\text{Be-F} = 632</math>, <math>\text{Be-I} = 289</math>, <math>\text{Hg-F} = 268</math> and <math>\text{Hg-I} = 145</math>)</p>

# Gate Chemistry Question Paper

**N Colangelo**



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