

Gas Law Stoichiometry Worksheet

Directions: Use significant figures and units in the problems below.

$PV = nRT$

SHOW ALL WORK!!!

1. Given the following unbalanced chemical equation for the combination reaction of sodium metal and chlorine gas:



- a. What volume of chlorine gas, measured at STP, is necessary for the complete reaction of 4.81 g of sodium metal.

$4.81 \text{ g Na} \left| \frac{1 \text{ mol}}{23.0 \text{ g Na}} \right| \left| \frac{1 \text{ mol Cl}_2}{2 \text{ mol Na}} \right| \left| \frac{22.4 \text{ L}}{1 \text{ mol}} \right| = 2.34 \text{ L Cl}_2$

2.34 L Cl_2



- a. Balance the above reaction.
 b. What volume of oxygen gas at 25°C and 1.04 atm is needed for the complete combustion of 5.53 g of propane?

$n = x$
 $V = O_2(x)$

$T_1 = 25 + 273 = 298 \text{ K}$
 $P_1 = 1.05 \text{ atm}$

$V = \left(\frac{5.53 \text{ g C}_3\text{H}_8}{44 \text{ g}} \right) \left(\frac{1 \text{ mol C}_3\text{H}_8}{1 \text{ mol C}_3\text{H}_8} \right) \left(\frac{5 \text{ mol O}_2}{1 \text{ mol C}_3\text{H}_8} \right) = 0.628 \text{ moles}$
 $V = 14.6 \text{ L}$

14.6 L

3. Potassium permanganate, KMnO_4 , is produced commercially by oxidizing aqueous potassium manganate, K_2MnO_4 .



- a. Balance the above reaction.
 b. What volume of $\text{Cl}_2(g)$, measured at STP, is needed to produce 10.0 g of KMnO_4 ?

$V = \frac{nRT}{P}$

$10.0 \text{ g KMnO}_4 \left| \frac{1 \text{ mol}}{158 \text{ g}} \right| \left| \frac{1 \text{ mol Cl}_2}{2 \text{ mol}} \right| \left| \frac{22.4 \text{ L}}{1 \text{ mol Cl}_2} \right| = 0.709 \text{ L Cl}_2$

0.709 L

4. If water is added to magnesium nitride, ammonia gas is produced when the mixture is heated.



- a. Balance the above reaction.
 b. If 10.3 g of magnesium nitride is treated with 10.3 g of water, what volume of ammonia gas would be collected at 24°C and 752 mmHg?

$10.3 \text{ g Mg}_3\text{N}_2 \left| \frac{1 \text{ mol}}{100.9 \text{ g}} \right| \left| \frac{2 \text{ mol NH}_3}{1 \text{ mol}} \right| = 0.204 \text{ mol}$
 Lim Reagent

$10.3 \text{ g H}_2\text{O} \left| \frac{1 \text{ mol}}{18 \text{ g}} \right| \left| \frac{2 \text{ mol NH}_3}{3 \text{ mol}} \right| = 0.381 \text{ mol}$

$24 + 273 = 297 \text{ K}$
 $752 / 760 = 0.989 \text{ atm}$

5.03 L

$V = \frac{nRT}{P}$

$V = \left(\frac{0.204 \text{ mol}}{0.989 \text{ atm}} \right) \left(\frac{297 \text{ K}}{0.989 \text{ atm}} \right) = 5.03 \text{ L}$

Gas Stoichiometry Answer Key 13

Jianjun Gao



Gas Stoichiometry Answer Key 13:

Gas Stoichiometry Answer Key 13 Book Review: Unveiling the Magic of Language

In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has become apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Gas Stoichiometry Answer Key 13**," written by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound affect our existence. Throughout this critique, we shall delve to the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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