

Answers

$$1) \quad c_p = 0.0042 \text{ J/g}^\circ\text{C}$$

$$= 345.34\text{g}(250.00^\circ\text{C} - 35.0^\circ\text{C})0.09\text{J/g}^\circ\text{C}$$

from there its just simple Algebra.

Type 1

Step 1: First read the question and try to understand what they are asking you. Can you picture a piece of aluminum foil that is taken out of an oven. Imagine the aluminum losing heat to its surroundings until the temperature goes from 415.0°C to 22.0°C .

Step 2: Write the original formula.

$$q = m(\Delta T)C_p$$

Step 3: List the known and unknown factors. Looking at the units in the word problem will help you determine which is which.

$$q = ?$$

$$m = 23.984 \text{ g}$$

$$\Delta T = (415.0^\circ\text{C} - 22.0^\circ\text{C}) = 393.0^\circ\text{C} \quad (\text{remember, they asked for the change in temperature})$$

$$C_p = 0.902 \text{ J/g} \times ^\circ\text{C}$$

Step 4: Substitute your values into the formula

$$q = ?$$

$$m = 23.984 \text{ g}$$

$$\Delta T = (415.0^\circ\text{C} - 22.0^\circ\text{C}) = 393.0^\circ\text{C}$$

$$C_p = 0.902 \text{ J/g} \times ^\circ\text{C}$$

$$q = m(\Delta T)C_p$$

$$q = 23.984 \text{ g} \times 393.0^\circ\text{C} \times 0.902 \text{ J/g} \times ^\circ\text{C}$$

Step 5: Cross out units where possible, and solve for unknown.

$$q = 23.984 \text{ g} \times 393.0^\circ\text{C} \times 0.902 \text{ J/g} \times ^\circ\text{C}$$

$$q = 8501.992224 \text{ J}$$

Ii Heat Problems Answers

SA Adler



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