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Date: 1/19/24

## Student Exploration: Bohr Model of Hydrogen

**Vocabulary:** absorption spectrum, Bohr model, electron volt, emission spectrum, energy level, ionization energy, laser, orbital, photon

*[Note to teachers and students: This Gizmo was designed as a follow-up to the Bohr Model: Introduction Gizmo. We recommend doing that activity before trying this one.]*

**Prior Knowledge Questions** (Do these BEFORE using the Gizmo.)

- What happens to an electron when it absorbs a tiny packet of light called a **photon**?  
the electron absorbs the energy from it, causing a higher energy level
- What happens to an electron when it emits a photon?  
the electrons lose energy, causing a lower energy level

### Gizmo Warm-up

When white light is passed through hydrogen gas and separated by a prism, some wavelengths of light are absorbed by the hydrogen atoms in the gas. This causes black bars to appear in the **absorption spectrum** of hydrogen. An **emission spectrum** is created when hydrogen gas emits light.

The **Bohr model** of the atom was inspired by the spectrum produced by hydrogen gas. The link between light and atomic structure is illustrated by the *Bohr Model of Hydrogen* Gizmo. The Gizmo shows a **laser** pointed at a container of hydrogen gas. After passing through the gas, light from the laser goes through a prism and is detected on a screen.

- With the **Laser energy** set to 7.0 eV, click **Play** (). Observe the blue electron on the ORBITALS pane. Do 7.0-eV photons have any effect on the electron? **No, it doesn't, the electron stays at lvl 1.**
- Click **Pause** (). Set the **Laser energy** to 12.1 eV and click **Play**. What happens to the electron now? **the electron shakes up and down and moves up by 2 orbital (level 3 )**
- Under **Go to energy level**, select **1**. Notice that a photon is emitted by the electron. What is the energy of the emitted photon? **(12.1 eV)**

# Explore Learning Exploration Guide Bohr Model Answers

**KJ Lindholm-Leary**



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## **Table of Contents Explore Learning Exploration Guide Bohr Model Answers**

1. Understanding the eBook Explore Learning Exploration Guide Bohr Model Answers
  - The Rise of Digital Reading Explore Learning Exploration Guide Bohr Model Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Explore Learning Exploration Guide Bohr Model Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Explore Learning Exploration Guide Bohr Model Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Explore Learning Exploration Guide Bohr Model Answers
  - Personalized Recommendations
  - Explore Learning Exploration Guide Bohr Model Answers User Reviews and Ratings
  - Explore Learning Exploration Guide Bohr Model Answers and Bestseller Lists
5. Accessing Explore Learning Exploration Guide Bohr Model Answers Free and Paid eBooks
  - Explore Learning Exploration Guide Bohr Model Answers Public Domain eBooks
  - Explore Learning Exploration Guide Bohr Model Answers eBook Subscription Services
  - Explore Learning Exploration Guide Bohr Model Answers Budget-Friendly Options
6. Navigating Explore Learning Exploration Guide Bohr Model Answers eBook Formats

- ePub, PDF, MOBI, and More
  - Explore Learning Exploration Guide Bohr Model Answers Compatibility with Devices
  - Explore Learning Exploration Guide Bohr Model Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Explore Learning Exploration Guide Bohr Model Answers
    - Highlighting and Note-Taking Explore Learning Exploration Guide Bohr Model Answers
    - Interactive Elements Explore Learning Exploration Guide Bohr Model Answers
  8. Staying Engaged with Explore Learning Exploration Guide Bohr Model Answers
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Explore Learning Exploration Guide Bohr Model Answers
  9. Balancing eBooks and Physical Books Explore Learning Exploration Guide Bohr Model Answers
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Explore Learning Exploration Guide Bohr Model Answers
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Explore Learning Exploration Guide Bohr Model Answers
    - Setting Reading Goals Explore Learning Exploration Guide Bohr Model Answers
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Explore Learning Exploration Guide Bohr Model Answers
    - Fact-Checking eBook Content of Explore Learning Exploration Guide Bohr Model Answers
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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