



Student Exploration: Boyle's Law and Charles's Law

Vocabulary: absolute zero, Boyle's law, Charles's law, Gay-Lussac's law, Kelvin scale, pressure

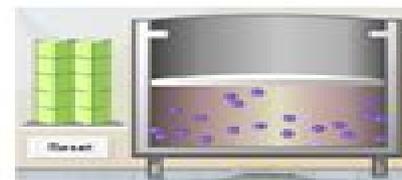
Prior Knowledge Question (Do this BEFORE using the Gizmo.)

A small helium tank measures about two feet (60 cm) high. Yet it can fill over 50 balloons! How can such a small tank contain enough helium to fill so many balloons?

The small container can hold all that helium in the container because the gas in the tank is under a lot of pressure.

Gizmo Warm-up

The *Boyle's Law and Charles's Law* Gizmo shows a container of gas. Inside, small purple spheres represent gas molecules.



1. Observe the particles. Are they all moving at the same speed? **Yes they are all moving at the same speed.**

2. How do the particles interact with the walls and lid of the container?

The bounce off the walls and lid while providing pressure

These interactions contribute to the **pressure** on the walls of the container. Pressure is defined as force per unit area. The SI units of pressure are newtons per square meter (N/m^2), or pascals (Pa).

3. Slowly drag the temperature (**T**) slider back and forth. (Note: In this Gizmo, the **Kelvin scale** is used to measure temperature. On the Kelvin scale, 0 degrees is **absolute zero**, the coldest possible temperature. Absolute zero is equal to -273.15°C or -459.67°F)

A. How does the change in temperature affect the speed of the molecules? Lower the temperature slower the molecules. Higher the temperature the faster the molecules.

How does the change in temperature affect the volume of the container? As the temperature gets higher the pressure increases and the lid goes higher due to the force being applied.



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Rachel S Tattersall



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