

IONIC BONDS

Ionic Bonds

Reading Passage

Read the passage carefully before answering the questions that follow.

Ionic bonds are a type of chemical bond formed due to the electrostatic attraction between oppositely charged ions. These ions are created when atoms transfer electrons, resulting in a cation (a positive ion) and a negative ion (anion). For instance, in sodium chloride (NaCl), sodium loses an electron to become a cation, while chlorine gains an electron to become an anion. The transfer of electrons creates ions that are attracted to each other because of their opposite charges.

The arrangement of ionic ions in compounds is often depicted as a crystal lattice – a highly ordered, repeating three-dimensional pattern. This structure maximizes the electrostatic attractions between oppositely charged ions, which leads to characteristic physical properties such as high melting points, brittleness, and good solubility in water. The ionic structure ensures that each ion is surrounded by ions of opposite charge, creating a stable configuration.

In the macroscopic level, the electrostatic attraction between ions results in a stable and highly ordered structure. This stability at the atomic level contributes to the distinctive properties of ionic compounds on a macroscopic scale. For example, the high melting points of ionic compounds are due to the significant energy required to overcome the strong electrostatic forces holding ions in the lattice structure. Similarly, ionic compounds are typically brittle because they tend to shatter rather than deform under stress, which can be attributed to repulsive interactions, leading to fracture.

True or False? Ionic Bond Concepts

Circle whether each statement about ionic bonds is True or False.

1. Ionic bonds form due to the electrostatic attraction between ions with opposite charges.	10
2. Chlorine atoms within gases are attracted to become a negative ion.	10
3. Ionic compounds are usually crystalline and have high melting points.	10
4. The arrangement of ions in compounds maximizes attraction between similar charges.	10
5. Electrostatic forces are the driving of reactions between atoms.	10
6. In the lattice structure, each ion is surrounded by ions of opposite charge.	10

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