

Name: _____

Date: _____

Student Exploration: Electron Configuration

Activity C:
The diagonal rule

Get the Gizmo ready:

- On the PERIODIC TABLE tab, select **Ar** (argon).
- Select the ELECTRON CONFIGURATION tab.
- Turn on **Show number of electrons**.

1s	↑↓		
2s	↑↓	2p	↑↓ ↑↓ ↑↓
3s	↑↓	3p	↑↓ ↑↓ ↑↓
4s	□	4p	□ □ □

Introduction: Beyond argon, it is a bit tricky to determine which subshell gets filled next. There are several rules that scientists use to determine the electron configurations of larger atoms.

Question: How are the electron configurations of elements beyond argon determined?

1. **Arrange:** Create the correct electron configuration for argon. Then, click **Next element** to get to potassium (K). Click once in the first **3d** orbital, and then click **Check**.

What feedback is given? _____

2. **Rearrange:** As it happens, the 4s subshell is a lower-energy subshell than 3d, so it is filled first. Remove the electron from the **3d** orbital and place it in the **4s** orbital. Click **Check**. (Note: For simplicity, all but the outer shell electrons will disappear on the **Bohr Model**.)

Is this configuration correct? _____ What is the configuration? _____

3. **Arrange:** Click **Next element** and add an electron for calcium. Click **Check**.

What is the electron configuration for calcium? _____

4. **Arrange:** Click **Next element** and add an electron for scandium. Try different orbitals until you find the right one.

What is the electron configuration for scandium? _____

5. **Observe:** Scandium is the first element to contain electrons in the *d* subshell. How many orbitals does the *d* subshell have, and how many electrons can fit in the *d* subshell?

6. **Infer:** Select the PERIODIC TABLE tab. The middle section of the table is a chemical family called the transition metals. Why do you think this section is ten columns wide?



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Christian Drosten



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