

General Chemistry Homework Answer Key

Textbook: OpenStax Chemistry: Atoms First 2e

Chapter 3: Electronic Structure and Periodic Properties of Elements

TA Reference Answers with Step-by-Step Explanations

3.1 Electromagnetic Energy

- 2 – Define frequency, wavelength, and amplitude.

Answer:

Frequency (ν) is the number of wave cycles that pass a point per second, measured in hertz (Hz).

Wavelength (λ) is the distance between successive crests of a wave, measured in meters (m).

Amplitude is the height of the wave from the center line to the crest or trough, indicating intensity.

Stepwise Procedure:

1. Define each term conceptually.
2. Include appropriate units.
3. Clarify differences among the properties.

- 4 – What is the relationship between frequency and energy?

Answer:

The energy (E) of electromagnetic radiation is directly proportional to its frequency (ν): $E = h\nu$, where h is Planck's constant (6.626×10^{-34} J·s).

Stepwise Procedure:

1. State the energy-frequency equation.
2. Define h and ν .
3. Explain that higher frequency means higher energy.

3.2 The Bohr Model

- 8 – What are the key features of the Bohr model of the hydrogen atom?

Answer:

1. Electrons orbit the nucleus in fixed paths (energy levels).
2. Only certain orbits are allowed with quantized energies.
3. Electrons absorb or emit energy when they move between levels.

General Chemistry Homework Guide

Wolfgang Guggemos



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