

General Chemistry Semester 2 Study Guide

Mr. Pitzer - Fall 2019

1st Period: Monday, June 17th, 8:00am - 10:00am

2nd Period: Wednesday, June 19th, 11:00am - 1:00pm

3rd Period: Thursday, June 20th, 11:00am - 1:00pm

Use your resources wisely throughout the semester. You may not get to the end of the course by a college of other students that are less active than themselves. Maximizing your learning means the use of resources that give us a chance to practice and extend a concept/idea. Please review your notes, HW, and Quizzes for additional insight.

Make yourself a learning schedule and the course content/assignments as a resource. Think of this as your study guide. You will suggest the content and when a big topic or lesson begins that will make your studying more efficient. Use the completed schedule/learning schedule you have used when that topic/lesson begins.

Our classes include an element of laboratory, experiments, assignments, and discussion.

Make sure to follow this:

Make sure you have the following resources necessary available and that you have access to your assignments throughout the semester.

- 1. Periodic Table
- 2. Chemical Properties Table
- 3. Molar Masses Table
- 4. Thermodynamic Tables
- 5. Learning your ability to recognize/analyze a chemical process through recognizing/understanding
 - 1. All substances below their standard for use above
 - o $\Delta H^{\circ}_{\text{f}}(\text{aq})$ and $\Delta G^{\circ}_{\text{f}}(\text{aq})$
 - o $\Delta H^{\circ}_{\text{f}}(\text{g})$
 - o $\Delta G^{\circ}_{\text{f}}(\text{g})$
- 6. Standard reduction potential table (electrochemistry)

Excellent Reading Challenge

Review each week's text for following sections and understand/analyze/evaluate/compare/contrast each of which all appropriate/possible.

- | | |
|--|---|
| 1. Molecular Structure Equations (10.1 - 10.4) <ul style="list-style-type: none">o Molecular structure and Lewis structureso VSEPRo Molecular geometry | 2. Thermodynamics (9.1 - 9.4) <ul style="list-style-type: none">o Heato Enthalpy |
| 3. Thermodynamics (9.5 - 9.7, 9.9 - 9.10) <ul style="list-style-type: none">o Heato Enthalpy | 4. Molecular Structure (10.1 - 10.4) <ul style="list-style-type: none">o Molecular structureo VSEPRo Molecular geometry |
| 5. Thermodynamics (9.1 - 9.4) <ul style="list-style-type: none">o Heato Enthalpy | 6. Kinetics and Reaction Equilibrium (14.1 - 14.5) <ul style="list-style-type: none">o Kineticso Reaction Equilibrium |
| 7. Equilibrium (9.1 - 9.4) <ul style="list-style-type: none">o Heato Enthalpy | |

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Table of Contents General Chemistry Semester 2 Study Guide

1. Understanding the eBook General Chemistry Semester 2 Study Guide
 - The Rise of Digital Reading General Chemistry Semester 2 Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying General Chemistry Semester 2 Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an General Chemistry Semester 2 Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from General Chemistry Semester 2 Study Guide
 - Personalized Recommendations
 - General Chemistry Semester 2 Study Guide User Reviews and Ratings
 - General Chemistry Semester 2 Study Guide and Bestseller Lists
5. Accessing General Chemistry Semester 2 Study Guide Free and Paid eBooks
 - General Chemistry Semester 2 Study Guide Public Domain eBooks
 - General Chemistry Semester 2 Study Guide eBook Subscription Services

- General Chemistry Semester 2 Study Guide Budget-Friendly Options
- 6. Navigating General Chemistry Semester 2 Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - General Chemistry Semester 2 Study Guide Compatibility with Devices
 - General Chemistry Semester 2 Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of General Chemistry Semester 2 Study Guide
 - Highlighting and Note-Taking General Chemistry Semester 2 Study Guide
 - Interactive Elements General Chemistry Semester 2 Study Guide
- 8. Staying Engaged with General Chemistry Semester 2 Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers General Chemistry Semester 2 Study Guide
- 9. Balancing eBooks and Physical Books General Chemistry Semester 2 Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection General Chemistry Semester 2 Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine General Chemistry Semester 2 Study Guide
 - Setting Reading Goals General Chemistry Semester 2 Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of General Chemistry Semester 2 Study Guide
 - Fact-Checking eBook Content of General Chemistry Semester 2 Study Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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