

ACS GENERAL CHEMISTRY FINAL STUDY GUIDE

Avagadro's Number - $N_A = 6.023 \times 10^{23}$

Percent Yield - $(\text{Actual Yield}/\text{Theoretical Yield}) \times 100$

Percent Composition (of an Element) - $[(\text{moles of element} \times \text{molar mass of element})/(\text{molar mass of compound})] \times 100$

Dilution Equation - $M_1V_1 = M_2V_2$
Mass and Volume

Boyle's Law - $P_1V_1 = P_2V_2$
Pressure and Volume

Charles' Law - $V_1/T_1 = V_2/T_2$
Volume and Temperature

Ideal Gas Law - $PV = nRT$

Volume of an Ideal Gas at STP - 22.4 L

Density (Ideal Gas) - $(P \times \text{Molar Mass})/(RT)$

Definition of Entahipy - $H = E + PV$

Definition of Heat Capacity - $C = ms$
 $s = \text{specific heat}$

ion - different number of electrons

isotopes - different number of neutrons

Periodic Table Trends - As you go down or left atomic radius increases
As you go up or right ionization energy and electron affinity increases

Ionic Bonds - to charged pieces come together and then there are cations or anions

Covalent Bonds - are when two non metals share electrons

A Compound is probably soluble if it contains these anions: - Cl^-
 Br^-
 I^-

General Chemistry Final Study Guide

Terry C. Jones



General Chemistry Final Study Guide:

Unveiling the Magic of Words: A Overview of "**General Chemistry Final Study Guide**"

In a world defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their power to kindle emotions, provoke contemplation, and ignite transformative change is really awe-inspiring. Enter the realm of "**General Chemistry Final Study Guide**," a mesmerizing literary masterpiece penned with a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve in to the book is central themes, examine its distinctive writing style, and assess its profound effect on the souls of its readers.

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