

## ARE YOU READY? PAGE 179

1. D
2. C
3. E
4. B
5. LCM =  $3^2 \times 2$   
= 18
6. LCM =  $2^3 \times 7$   
= 56
7. LCM =  $2^3 \times 3^2$   
= 72
8. LCM =  $5 \times 3^2$   
= 135
9. perpendicular
10. parallel;  
 $5x - 10y = 3$   
 $-10y = -5x + 3$   
 $y = \frac{1}{2}x - \frac{3}{10}$
- and  
 $y = \frac{1}{2}x - 6$
11. perpendicular;  
 $x - y = 3$   
 $y = x - 3$
- and  
 $x + y = 4$   
 $y = -x + 4$
- and  
 $3y - x = 5$   
 $3y = x + 5$   
 $y = \frac{1}{3}x + \frac{5}{3}$
12. neither;  
 $2x - 3y = -4$   
 $3y = 2x + 4$   
 $y = \frac{2}{3}x + \frac{4}{3}$
13.  $1.5(8) + 3(14)$   
=  $12 + 42$   
= 54
14.  $5(6) - \frac{3}{4}(-4)$   
=  $30 + 3$   
= 33
15.  $4(0.25) - \sqrt{2}(\sqrt{2})$   
=  $1 - \sqrt{4}$   
=  $1 - 2$   
= -1
16.  $-\frac{75(1)}{3(\frac{1}{3})}$   
=  $-\frac{75}{1}$   
= -75
17.  $6x + 19 = -5$   
 $6x = -24$   
 $x = -3$
18.  $5x + 4 = 25 - 2x$   
 $7x + 4 = 25$   
 $7x = 21$   
 $x = 3$
19.  $9x - (x + 12) = -13$   
 $9x - x - 12 = -13$   
 $8x - 12 = -13$   
 $8x = -1$   
 $x = -\frac{1}{8}$
20.  $-3(4x - 5) - 1 = 20$   
 $-12x + 15 - 1 = 20$   
 $-12x + 14 = 20$   
 $-12x = 6$   
 $x = -\frac{6}{12}$   
 $x = -\frac{1}{2}$
21.  $\frac{1}{4}x + \frac{2}{3}x = 6$   
 $3x + 8x = 96$   
 $11x = 96$   
 $x = \frac{96}{11}$
22.  $\frac{2}{5}x + \frac{1}{6} = -4$   
 $12x + 5 = -120$   
 $12x = -125$   
 $x = -\frac{125}{12}$

$$23. \begin{array}{r} x + \frac{1}{2} = -\frac{1}{9} \\ 10x + 5 = -2 \\ 10x = -7 \\ x = -\frac{7}{10} \end{array}$$

$$24. \begin{array}{r} -\frac{1}{2} = 3x - \frac{1}{3}x \\ -3 = 18x - 2x \\ -3 = 16x \\ -\frac{3}{16} = x \end{array}$$

## 3-1 USING GRAPHS AND TABLES TO SOLVE LINEAR SYSTEMS, PAGES 182–189

## CHECK IT OUT!

1a.  $\begin{array}{r} x + 2y = 10 \\ (4) + 2(3) \quad | \quad 10 \\ 10 \quad | \quad 10 \checkmark \end{array}$   
(4, 3) is a solution.

$$\begin{array}{r} 3x - y = 9 \\ 3(4) - (3) \quad | \quad 9 \\ 9 \quad | \quad 9 \checkmark \end{array}$$

b.  $\begin{array}{r} 6x - 7y = 1 \\ 6(5) - 7(3) \quad | \quad 1 \\ 9 \quad | \quad 1 \neq \end{array}$   
(5, 3) is not a solution.

2a.  $\begin{cases} 2y + 6 = x \\ 4x = 3 + y \end{cases}$   
 $y = \frac{1}{2}x - 3$

x	y
2	-2
0	-3
-2	-4

$$y = 4x - 3$$

x	y
2	5
0	-3
-2	-11

The solution to the system is (0, -3).

b.  $\begin{cases} x + y = 8 \\ 2x - y = 4 \\ y = -x + 8 \end{cases}$

x	y
0	8
2	6
4	4

$$y = 2x - 4$$

x	y
0	-4
2	0
4	4

The solution to the system is (4, 4).

c.  $\begin{cases} y - x = 5 \\ 3x + y = 1 \end{cases}$

$$y = x + 5$$

x	y
-1	4
0	5
1	6

$$y = -3x + 1$$

x	y
-1	4
0	1
1	-2

The solution to the system is (-1, 4).

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**N Colangelo**



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web jan 1 2012 in experiment 19 the theory of atomic spectra we study the emission lines from the hydrogen atom in terms of the bohr theory of the atom the light coming from a hydrogen discharge tube is spread into its constituent wavelengths by a diffraction grating

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web experiment sweeping the detector arm through will now record a spectrum of the light from the discharge lamp try different aperture and slit sizes and adjusting the lens locations to record as many of the spectral peaks as possible dim spectral peaks will require careful tuning of the apertures to observe

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web the atomic spectra experiment is a complete solution for measuring the wavelengths of the discrete lines of the atomic spectra of various gases using a grating spectrophotometer providing educators worldwide with innovative solutions for teaching science

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web to observe and understand line emission spectra of atoms using gas discharge tubes to practice writing electron configurations for these and other elements electromagnetic radiation is energy in the form of waves waves are characterized by their wavelength  $\lambda$  and frequency  $\nu$

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web what is happening within an atom that causes it to emit light in specific lines in a spectrum why did the different elements have different line emission spectra the radio station jack fm broadcasts at a frequency of 93.5 khz

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web the field of atomic spectra provides in fact a perfect example of the way in which experiment and theory can react

fruitfully on each other yielding advances in our understanding of the

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web atomic spectroscopy experiment objectives test a diffraction grating based spectrometer study the energy spectrum of atomic hydrogen h and a hydrogen like atomic sodium na determine values of quantum defects of low angular momentum states of na and measure fine splitting using na yellow doublet history

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web the bohr model named after danish physicist niels bohr of an atom has a small positively charged central nucleus and electrons orbiting in at specific fixed distances from the nucleus

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web objectives to perform flame tests of metal cations in order to observe their featured colors to perform calculations up determine the frequency and energy in the emitted photons to tell these results the the types of electronic transitions occurring in are default

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web atomic emission spectra electrons in atoms normally occupy the lowest energy states possible such an atom is said to be in the ground state however electrons can be excited to high energy states when they absorb excess energy the excess energy can be provided by heat light or electrical discharge

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web atomic spectra experiment introduction in this experiment you will observe the spectra lines emitted by hydrogen gas and mercury gas contained in spectral tubes equipment optical bench meter stick diffraction grating lamp power supply hydrogen and mercury spectral tubes theory

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web specialists on hydrogen and hydrogen like systems presents the advances in theory and experiment that have occurred since that time behavior of electrons in atoms nov 30 2022 atomic spectra the interactions between atoms and electrons quantum theory of atomic structure

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web lab 7 atomic spectra chemistry lab notes general chemistry i chem 1411 students shared 265 documents in this course more info the purpose of the lab was to determine the characteristics of visible light that was emitted by different elements w e did this by conducting a flame test the purpose was

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web fingerprints of the elements atomic spectra as dr matilsky discussed in his video lecture atomic spectra occur due to the

fact that orbital radii of electrons and hence their energies are quantized at specific levels determined by the atomic number number of protons and ionization state number of electrons in any given element

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web in this experiment the student measures the features of a simple spectrum that of sodium an empirical relation is deduced for the frequencies of the lines and the position of unobserved lines is predicted