

Intermolecular Forces HOMEWORK

Lesson 1

1. The Lewis Dot Structure of Boron consists of one PAIR of electrons and one lone electron. This suggests that Boron only has one free electron to form a bond. So, it seems logical that Boron should only be able to bond to one atom. How is it possible for Boron to form a compound such as BCl_3 ? Use valence bond theory to explain this.

Clue w/ hybridization: There are seven sp^2 hybrid orbitals and a hybrid orbital with one of the electrons are paired. That one Boron will form 3, unpaired electrons and be able to bond to three other atoms. (Boron has 3 paired electrons in the s orbital and only 1 unpaired electron in the p orbital)

2. What type of hybridized orbital is used by the central atom of each of these molecules? (You will need to refer to their Lewis dot structures.)



3. True or False? This triple bond in N_2 results from two sp sigma bonds and one sp pi bond. **False** sp sp sp
 (The triple bond is made of 1 sigma bond and 2 pi bonds.)

4. These molecules are shown below. Are these three molecules isomers? **Yes, because they all have the formula C_2H_4O**



5. Indicate the hybridization of each C atom in each molecule.
 vinyl alcohol sp^2 ethylene oxide sp^3 acetylenediol sp left carbon sp^2 right carbon sp^3

6. How many pi bonds are in each of the molecules above?
 vinyl alcohol 1 ethylene oxide 0 acetylenediol 2

7. What is the hybridization of the carbon atom in $CaCO_3$ sp^2
 a) How many sigma bonds are in $CaCO_3$ 3
 b) How many pi bonds are in $CaCO_3$ 1



8. Ammonia can form a covalent bond with a hydrogen ion as shown below.



- What is the name of the specific type of covalent bond that forms? **coordinate covalent bond**
- What is the hybridization of the nitrogen atom in ammonia before it bonds? sp^3
- What is the hybridization of the nitrogen atom in the ammonium ion? sp^3
- Does the nitrogen atom's hybridization change when it forms this special type of covalent bond? **No**

9. Benzene, a resonance structure that is commonly used to make plastics, is a covalent compound that conducts electricity. Most compounds do NOT conduct electricity, since they do not contain any ions. Use the molecular orbital theory to explain this unexpected conductivity.
Carbon's valence orbitals do not combine ions, but we can transfer the lone paired pair of electrons into delocalized and spread throughout the molecule. This separation allows them to freely move, and therefore conduct electricity.

Holt Chemistry Answer Key Section Intermolecular Forces

Ralph H. Petrucci, William S. Harwood

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Holt Chemistry Ralph Thomas Myers, 2004 *Principles of Modern Chemistry* David W. Oxtoby, H. Pat Gillis, Laurie J. Butler, 2016-01-01 Long considered the standard for honors and high level mainstream general chemistry courses *PRINCIPLES OF MODERN CHEMISTRY* continues to set the standard as the most modern rigorous and chemically and mathematically accurate text on the market This authoritative text features an atoms first approach and thoroughly revised chapters on Quantum Mechanics and Molecular Structure Chapter 6 Electrochemistry Chapter 17 and Molecular Spectroscopy and Photochemistry Chapter 20 In addition the text utilizes mathematically accurate and artistic atomic and molecular orbital art and is student friendly without compromising its rigor End of chapter study aids focus on only the most important key objectives equations and concepts making it easier for students to locate chapter content while applications to a wide range of disciplines such as biology chemical engineering biochemistry and medicine deepen students understanding of the relevance of chemistry beyond the classroom **Host-Guest Chemistry** Brian D. Wagner, 2020-10-26 This textbook addresses the chemical and physicochemical principles of supramolecular host guest chemistry in solution It covers the thermodynamics and dynamics of inclusion and highlights several types of organic hosts Various applications of host guest chemistry in analytical and environmental chemistry as well as pharmaceutical and chemical industry demonstrate the versatile usability of molecular cages *Molecular Driving Forces* Ken Dill, Sarina Bromberg, 2010-10-21 *Molecular Driving Forces* Second Edition E book is an introductory statistical thermodynamics text that describes the principles and forces that drive chemical and biological processes It demonstrates how the complex behaviors of molecules can result from a few simple physical processes and how simple models provide surprisingly accurate insights into the workings of the molecular world Widely adopted in its First Edition *Molecular Driving Forces* is regarded by teachers and students as an accessible textbook that illuminates underlying principles and concepts The Second Edition includes two brand new chapters 1 Microscopic Dynamics introduces single molecule experiments and 2 Molecular Machines considers how nanoscale machines and engines work The Logic of Thermodynamics has been expanded to its own chapter and now covers heat work processes pathways and cycles New practical applications examples and end of chapter questions are integrated throughout the revised and updated text exploring topics in biology environmental and energy science and nanotechnology Written in a clear and reader friendly style the book provides an excellent introduction to the subject for novices while remaining a valuable resource for experts [Class 11-12 Chemistry MCQ \(Multiple Choice Questions\)](#) Arshad Iqbal, 2019-05-17 The Class 11 12 Chemistry Multiple Choice Questions MCQ Quiz with Answers PDF College Chemistry MCQ PDF Download Quiz Questions Chapter 1 6 Practice Tests with Answer Key 11th 12th Grade Chemistry Questions Bank MCQs Notes includes revision guide for problem solving with hundreds of solved MCQs Class 11 12 Chemistry MCQ with Answers PDF book covers basic concepts analytical and practical assessment tests Class 11 12 Chemistry MCQ PDF book helps to practice test questions

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