

The curve in Figure 16.11 of the text shows how the vibrational heat capacity depends on temperature. Note that even when the temperature is only slightly above the vibrational temperature, the heat capacity is close to its equipartition value.

The total heat capacity of a molecular substance is the sum of each contribution (Figure 16.12 of the text). When equipartition is valid (when the temperature is well above the characteristic temperature of the mode $T \gg \theta_M$) we can estimate the heat capacity by counting the numbers of modes that are active. In gases, all three translational modes are always active and contribute $\frac{3}{2}R$ to the molar heat capacity. If we denote the number of active rotational modes by ν_R^\ddagger (so for most molecules at normal temperatures $\nu_R^\ddagger = 2$ for linear molecules, and 3 for non-linear molecules), then

$$C_{V,m} = \frac{3}{2}R + \frac{\nu_R^\ddagger}{2}R + \dots$$

The equipartition value $C_{V,m} = 2.5R$ is reached when all the vibrational modes are active. At $T = 298 \text{ K}$, the vibrational temperatures of the modes of CO_2 are 3300 K , 2200 K , 1330 K , 690 K , 470 K , and 430 K . The vibrational temperatures of the modes of H_2O are 5400 K , 3650 K , 1590 K , 940 K , 540 K , and 430 K . The vibrational temperatures of the modes of CH_4 are 4300 K , 2900 K , 1460 K , 950 K , 550 K , and 430 K . The vibrational temperatures of the modes of NH_3 are 4300 K , 2900 K , 1460 K , 950 K , 550 K , and 430 K .

Solutions to exercises

The vibrational energy is given by $\epsilon = \frac{1}{2}h\nu$ and the vibrational temperature is $\theta_v = \frac{h\nu}{k}$. The vibrational energy is $\epsilon = \frac{1}{2}k\theta_v$.

$$\epsilon = \frac{1}{2}k\theta_v = \frac{1}{2}(1.38 \times 10^{-23} \text{ J K}^{-1})(298 \text{ K}) = 2.04 \times 10^{-21} \text{ J}$$

The vibrational energy is $\epsilon = \frac{1}{2}k\theta_v = \frac{1}{2}(1.38 \times 10^{-23} \text{ J K}^{-1})(298 \text{ K}) = 2.04 \times 10^{-21} \text{ J}$.

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