

Study Guide: Concept 1 - Introduction to Bonding

Key Topics and Objectives

1. Forces Between Substances vs. Within a Substance

- **Intermolecular forces:** Forces between molecules (e.g., hydrogen bonding, dipole-dipole, London dispersion).
- **Intramolecular forces:** Forces within a molecule (e.g., ionic, covalent, and metallic bonds).

2. Bonding as a Spectrum

- Bonding is not just "ionic" or "covalent" but exists on a **spectrum** depending on electronegativity differences.
- **Pure covalent (nonpolar) → Polar covalent → Ionic**

3. Chemical Bonds, Reactions, and Compounds

- **Chemical bonds:** Forces holding atoms together.
- **Chemical reactions:** Process where substances change into new substances.
- **Compounds:** Substances formed when two or more elements chemically bond.

4. Why Atoms Form Chemical Bonds

- Atoms bond to **achieve stability** (noble gas electron configuration).
- The **Octet Rule** states that atoms gain, lose, or share electrons to reach **eight valence electrons**.

5. Chemical Formulas: Ionic vs. Covalent

- **Ionic compounds:** Metal + Nonmetal (transfer of electrons).
- **Covalent compounds:** Nonmetal + Nonmetal (sharing of electrons).

6. Covalent vs. Ionic Bonds

| Property | Covalent Bonds | Ionic Bonds |
|--------------------|----------------|--------------------|
| Electron Transfer? | Shared | Transferred |
| Types of Elements | Nonmetals | Metals & Nonmetals |

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Rachel Sandford



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