

Like Dissolves Like Practice Problems

Using your knowledge of the "like dissolves like" principle of polarity, determine which component will dissolve best in each of the following solvents. Explain your choice for the component that does not dissolve.

1) Solvent: methanol (CH₃OH)

Components: carbon tetrachloride and oxygen dichloride

Both oxygen dichloride and methanol are polar, while carbon tetrachloride is not.

2) Solvent: carbon dioxide

Components: methanol (CH₃OH) and nitrogen dioxide

While both components are polar, methanol is considerably more polar than nitrogen dioxide (both components always have that charge separation, while nitrogen dioxide, on certain times prefer, the not).

3) Solvent: water

Components: carbon tetrachloride and ammonia

Both water and ammonia are polar. Carbon tetrachloride is not.

4) Solvent: hexane (C₆H₁₄)

Components: water and methanol

Both water and methanol are polar. Methanol is not.

Like Dissolves Like Practice Answer Key

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