

## Firefighter 1, Chapter 2 (IFSTA Essentials of Firefighting, 6th ed.)

Strains and bruises - ANSWER What are the two main types of injuries that firefighters encounter?

Maintain high levels of physical fitness - ANSWER What is the best way to prevent common firefighter injuries?

Cardiovascular disease - ANSWER What is the leading cause of firefighter fatalities?

Requiring SCBA to be worn - ANSWER What is a departmental responsibility in the prevention of respiratory diseases?

Diabetes - ANSWER What illness can be signaled by high levels of sugar in the blood?

Use correct protective clothing when treating patients - ANSWER What is the proper way to prevent exposure-related diseases?

1851 - ANSWER What NFPA standard sets requirements for care and maintenance of personal and respiratory protection equipment?

Safe training practices - ANSWER Both NFPA 1407 and 1451 define

1582 - ANSWER What NFPA standard establishes safety programs departments must adopt?

Drug abuse - ANSWER A safety program established by NFPA 1500 must include provisions for what non-emergency issue?

Obvious fire fighting hazards - ANSWER What best describes what must be addressed in a department's safety and health program?

Training must meet NFPA 1000 ProQual standards - ANSWER What best describes NFPA 1500 requirements for training and education?

Meet the current edition of NFPA design standards - ANSWER NFPA 1500 requires the protective equipment must:

Have space and means for storing infection control devices - ANSWER According to NFPA 1500 fire facilities must:

# Essentials Of Firefighting 6th Edition Quizzes

**Hongru Du**



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When the half reactions listed above are combined into a balanced chemical equation the ratio of the coefficients of  $\text{Cu}$ ,  $\text{S}$ ,  $\text{MnO}_4^-$ ,  $\text{H}_2\text{O}$ , and  $\text{H}^+$  is

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web solution 1 the two half reactions balanced as if in acidic solution  $2\text{CN}^- + \text{Au} + \text{CN}^- + 2\text{e}^- + 2\text{H}^+ + \text{O}_2 \rightarrow \text{H}_2\text{O}_2 + 2\text{C}_2\text{O}_4^{2-}$  make electrons equal convert to basic solution  $4\text{CN}^- + 2\text{Au} + 2\text{CN}^- + 2\text{e}^-$  multiplied by a factor of 2  $2\text{e}^- + 2\text{H}_2\text{O} + \text{O}_2 \rightarrow \text{H}_2\text{O}_2 + 2\text{OH}^-$  the final answer

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation ni s ni a 2 a q 2 e now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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as no3 io3 re io3 h2aso4

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