

Intermolecular Forces HOMEWORK

Lesson 1

1. The Lewis Dot Structure of Boron consists of one PAIR of electrons and one lone electron. This suggests that Boron only has one free electron to form a bond. So, it seems logical that Boron should only be able to bond to one atom. How is it possible for Boron to form a compound such as BCl_3 ? Use valence bond theory to explain this.

Clue w/ hybridization: There are seven sp^2 hybrid orbitals and a hybrid orbital with one of the electrons are paired. That one orbital will form 3 sp^2 hybrid orbitals and be able to bond to three other atoms. (However, Boron has 3 paired electrons in the s orbital and only 1 unpaired electron in the p orbital)

2. What type of hybridized orbital is used by the central atom of each of these molecules? (You will need to refer to their Lewis dot structures.)



3. True or ~~False~~ This triple bond in N_2 results from two sp sigma bonds and one sp pi bond. **$:N \equiv N:$**
 (Example bond is made of 1 sigma bond and 2 pi bonds)

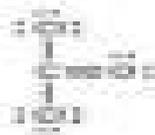
4. These molecules are shown below. Are these three molecules isomers? **Yes, because they all have the formula C_2H_4O**



5. Indicate the hybridization of each C atom in each molecule:
 vinyl alcohol sp^2 ethylene oxide sp^3 acetylene oxide sp left carbon sp^2 right carbon sp^3

6. How many pi bonds are in each of the molecules above?
 vinyl alcohol 1 ethylene oxide 0 acetylene oxide 2

7. What is the hybridization of the carbon atom in $CaCO_3$ sp^2
 a) How many sigma bonds are in $CaCO_3$ 3
 b) How many pi bonds are in $CaCO_3$ 2



8. Ammonia can form a covalent bond with a hydrogen ion as shown below.



- What is the name of the specific type of covalent bond that forms? **coordinate covalent bond**
- What is the hybridization of the nitrogen atom in ammonia before it bonds? sp^3
- What is the hybridization of the nitrogen atom in the ammonium ion? sp^3
- Does the nitrogen atom's hybridization change when it forms this special type of covalent bond? **No**

9. Benzene, a resonance structure that is commonly used to make plastics, is a covalent compound that conducts electricity. Most compounds do NOT conduct electricity, since they do not contain any ions. Use the molecular orbital theory to explain this unexpected conductivity.

Carbon's valence sp^2 orbitals form π bonds, but the unhybridized $2p_z$ orbitals are delocalized and spread throughout the molecule. This delocalization allows them to freely move, and therefore conduct, electrons.

Holt Chemistry Answer Key Section Intermolecular Forces

Arturo Cuomo



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Unveiling the Magic of Words: A Review of "**Holt Chemistry Answer Key Section Intermolecular Forces**"

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