

## Chapter 1: Solutions to Review Problems

### Section 1.1

- 1) If the number of moles of gas and the temperature remain constant:  $P_1V_1 = P_2V_2$

Substitute:  $100 \text{ kPa} \times 2.00 \text{ L} = P_2 \times 3.00 \text{ L} \therefore P_2 = 66.7 \text{ kPa}$

- 2) If the number of moles of gas and the pressure remain constant:  $\frac{V_1}{T_1} = \frac{V_2}{T_2}$

Substitute:  $\frac{1.00 \text{ L}}{200 \text{ K}} = \frac{V_2}{400 \text{ K}} \therefore V_2 = 2.00 \text{ L}$

- 3) If the number of moles of gas and the volume remain constant:  $\frac{P_1}{T_1} = \frac{P_2}{T_2}$

Substitute:  $\frac{100 \text{ kPa}}{298 \text{ K}} = \frac{P_2}{398 \text{ K}} \therefore P_2 = 133.6 \text{ kPa}$

- 4) Xe has molar mass  $131.3 \text{ gmol}^{-1}$ , Ar has molar mass  $39.95 \text{ gmol}^{-1}$ , Ne has molar mass  $20.18 \text{ gmol}^{-1}$

If equal masses of each gas are used, the sample of the gas with the smallest molar mass will contain the greatest number of moles.

Therefore, the sample of Ne will contain the greatest number of moles.

Pressure is directly proportional to the number of moles of gas; so the pressure in the Ne flask is the greatest. (Answer a)

- 5)  $\text{O}_2$  has molar mass  $32.00 \text{ gmol}^{-1}$ ,  $\text{N}_2$  has molar mass  $28.02 \text{ gmol}^{-1}$

If equal masses of each gas are used, the sample of the gas with the smaller molar mass will contain the greater number of moles.

Therefore, the sample of  $\text{N}_2$  will contain the greater number of moles, and therefore molecules.

Pressure is directly proportional to the number of moles (or molecules) of gas; so the pressure in the  $\text{N}_2$  flask is the greater than the pressure in the  $\text{O}_2$  flask. (Answer c)

# Holt Chemistry Chapter 1 solutions Test

**Ying Liu**



### **Holt Chemistry Chapter 1 solutions Test:**

**Holt Chemistry File** ,1998 This reference is a must for students who need extra help reteaching or extra practice The guide moves students through the same concepts as the text but at a slower pace More descriptive detail along with visual algorithms provides a more structured approach Each chapter closes with a large bank of practice problems Book jacket

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web you can use the post lab on the third page as a formative assessment where students label a diagram and match the parts with their roles this lab is aligned with florida s

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web activity begin the activity by reviewing the male and female parts of a flower it s helpful to have a diagram on the board you will want to review the following points to reinforce

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web directions for dissection on the back side of this sheet fill in a box with the name how many and function of each part color code the box to match the flower part 1 find and

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