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### Mixed Gas Law Calculations (Boyle's, Charles', and Gay-Lussac's)

Instructions: Solve each of the following problems **and** identify the law.

- 1) 2.00 L of a gas is at 740.0 mmHg pressure. What is its volume at standard pressure?

Law: Boyles

(Ans: 1.95 L)

$$P_1 \cdot V_1 = P_2 \cdot V_2$$

$$740.0 \cdot 2.00 = 101.325 \cdot V$$

$$V = 1.95$$

std pressure:  
1 atm = 101.325 mmHg

- 2) A gas is collected and found to fill 2.85 L at 25.0°C. What will be its volume at standard temp? Law: Charles

(Ans: 2.61 L)

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

$$\frac{2.85 \text{ L}}{298.15 \text{ K}} = \frac{V}{273.15 \text{ K}} = 3.1$$

273.15 K  
standard temp

- 3) A chemist collected 56.1 mL of gas in an open manometer. The next day, the chemist noted that the volume had changed to 57.9 mL and the barometer reading was 99.4 kPa. The temperature had not changed. What had been the barometer reading on the previous day when the gas was collected?

Law: Boyles

(Ans: 1.01 atm)

$$P_1 \cdot V = P_2 \cdot V$$

$$56.1 \text{ mL} \cdot P = 99.4 \text{ kPa} \cdot 57.9 \text{ mL} = 102.6$$

- 4) A glass sphere is filled to full volume with a gas. The pressure of the gas inside the sphere is 30.0 atm, and the temperature is 25.0°C. The sphere is taken outside on a cold day. The temperature of the gas decreases to 10.0°C. What is the new pressure of the gas? Assume that the volume is constant.

Law: G-L

(Ans: 28.5 atm)

$$\frac{P}{T} = \frac{P}{T}$$

$$\frac{30.0 \text{ atm}}{298^\circ\text{C}} = \frac{P}{10^\circ\text{C}} = 12$$

- 5) The gas pressure in an aerosol canister is 151.6 kPa at 25.0°C, what would the pressure be inside the canister at 300.0°C?

Law: G-L

(Ans: 2.88 atm)

$$\frac{P}{T} = \frac{P}{T}$$

$$\frac{151.6 \text{ kPa}}{298} = \frac{P}{300} = 1519.2$$

# Mixed Gas Law Calculation Work For Problems

**Valery Ochkov**



## **Mixed Gas Law Calculation Work For Problems:**

CliffsNotes Chemistry Practice Pack Charles Henrickson, 2010-02-08 About the Contents Pretest Helps you pinpoint where you need the most help Topic Area Reviews Measurement and Units of Measurement Matter Elements Compounds and Mixtures Atoms I The Basics Formulas and Names of Ionic Compounds Acids and Bases The Mole Elements and Compounds Percent Composition and Empirical and Molecular Formulas Chemical Reactions and Chemical Equations Calculations Using Balanced Equations Atoms II Atomic Structure and Periodic Properties Chemical Bonding The Formation of Compounds Gases and the Gas Laws The Forces between Molecules Solids and Liquids Solutions and Solution Composition Acids Bases and Neutralization Glossary Customized Full Length Exam Covers all subject areas Pretest that pinpoints what you need to study most Clear concise reviews of every topic Targeted example problems in every chapter with solutions and explanations Customized full length exam that adapts to your skill level **US Navy diving manual**, 1981

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