

Chemistry – Unit 5 Worksheet 1

Atomic Structure of the Elements

Exercise 1: Mass From Counts

By a better understanding of atomic structure, we are able to understand the relationship between mass and structure.

1. Equal numbers of atoms contain the same number of nucleons as the same mass of an element. **Simplest Equation**
2. $\text{Number of atoms} \times \text{mass of one atom} = \text{total mass of element}$

Ex. 1 10^6 ^{12}C atoms = 12g. 10^6 ^{16}O atoms = 16g. The density of oxygen gas is 1.43 g/litre and pressure is 1.01 g/cm³. Calculate the mass of oxygen gas in 1 litre of oxygen at 1.01 g/cm³. **Mass of 1 litre of oxygen = 1.43 g**

10⁶ atoms
12g ^{12}C
1 litre

10⁶ atoms
16g ^{16}O
1 litre



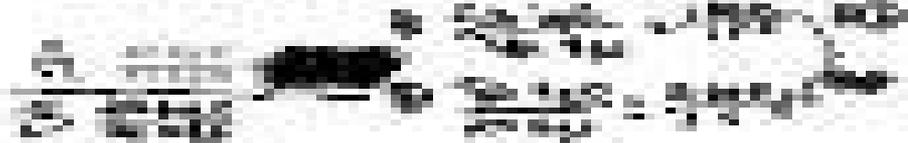
10⁶ atoms
1.43g O_2
1 litre
1.43g O_2
1.43g O_2

Calculate the mass of 10⁶ atoms of oxygen gas. The mass of 10⁶ atoms of oxygen gas is 16g. The mass of 10⁶ atoms of oxygen gas is 16g. The mass of 10⁶ atoms of oxygen gas is 16g.

Exercise 2: Mass From Counts

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R Sandford



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