

# Physics

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(Chapter 2) (Units and Measurement) (Exercises)

(Class 11)

## Question 2.15:

A famous relation in physics relates 'moving mass'  $m$  to the 'rest mass'  $m_0$  of a particle in terms of its speed  $v$  and the speed of light,  $c$ . (This relation first arose as a consequence of special relativity due to Albert Einstein). A boy recalls the relation almost correctly but forgets where to put the constant  $c$ . He writes:

$$m = \frac{m_0}{(1 - v^2)^{1/2}}$$

Guess where to put the missing  $c$ .

## Answer 2.15:

Given the relation,

$$m = \frac{m_0}{(1 - v^2)^{1/2}}$$

Dimension of  $m = M^1 L^0 T^0$ ,

Dimension of  $v = M^0 L^1 T^{-1}$ ,

Dimension of  $c = M^0 L^1 T^{-1}$

Dimension of  $m_0 = M^1 L^0 T^0$

Dimension of  $v^2 = M^0 L^2 T^{-2}$

The given formula will be dimensionally correct only when the dimension of L.H.S is the same as that of R.H.S.

This is only possible when the factor,  $(1 - v^2)^{1/2}$  is dimensionless i.e.,  $(1 - v^2)$  is dimensionless.

This is only possible if  $v^2$  is divided by  $c^2$ .

Hence, the correct relation is

$$m = \frac{m_0}{\left(1 - \frac{v^2}{c^2}\right)^{1/2}}$$

## Question 2.16:

The unit of length convenient on the atomic scale is known as an angstrom and is denoted by  $\text{\AA}$ :  $1 \text{\AA} = 10^{-10} \text{ m}$ . The size of a hydrogen atom is about  $0.5 \text{\AA}$ . What is the total atomic volume in  $\text{m}^3$  of a mole of hydrogen atoms?

## Answer 2.16:

Radius of hydrogen atom,  $r = 0.5 \text{\AA} = 0.5 \times 10^{-10} \text{ m}$

Volume of hydrogen atom =  $\frac{4}{3} \pi r^3$

$$= \frac{4}{3} \times \frac{22}{7} \times (0.5 \times 10^{-10})^3 = 0.524 \times 10^{-30} \text{ m}^3$$

1 mole of hydrogen contains  $6.023 \times 10^{23}$  hydrogen atoms.

Volume of 1 mole of hydrogen atoms =  $6.023 \times 10^{23} \times 0.524 \times 10^{-30}$

$$= 3.16 \times 10^{-7} \text{ m}^3$$

## Question 2.17:

One mole of an ideal gas at standard temperature and pressure occupies  $22.4 \text{ L}$  (molar volume). What is the ratio of molar volume to the atomic volume of a mole of hydrogen? (Take the size of hydrogen molecule to be about  $1 \text{\AA}$ ). Why is this ratio so large?

## Answer 2.17:

Radius of hydrogen atom,  $r = 0.5 \text{\AA} = 0.5 \times 10^{-10} \text{ m}$

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## Table of Contents Grade 11 Physcs Paper 2

1. Understanding the eBook Grade 11 Physcs Paper 2
  - The Rise of Digital Reading Grade 11 Physcs Paper 2
  - Advantages of eBooks Over Traditional Books
2. Identifying Grade 11 Physcs Paper 2
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Grade 11 Physcs Paper 2
  - User-Friendly Interface
4. Exploring eBook Recommendations from Grade 11 Physcs Paper 2
  - Personalized Recommendations
  - Grade 11 Physcs Paper 2 User Reviews and Ratings
  - Grade 11 Physcs Paper 2 and Bestseller Lists
5. Accessing Grade 11 Physcs Paper 2 Free and Paid eBooks
  - Grade 11 Physcs Paper 2 Public Domain eBooks
  - Grade 11 Physcs Paper 2 eBook Subscription Services
  - Grade 11 Physcs Paper 2 Budget-Friendly Options
6. Navigating Grade 11 Physcs Paper 2 eBook Formats
  - ePub, PDF, MOBI, and More
  - Grade 11 Physcs Paper 2 Compatibility with Devices
  - Grade 11 Physcs Paper 2 Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Grade 11 Physcs Paper 2
  - Highlighting and Note-Taking Grade 11 Physcs Paper 2
  - Interactive Elements Grade 11 Physcs Paper 2
8. Staying Engaged with Grade 11 Physcs Paper 2

- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Grade 11 Physcs Paper 2
9. Balancing eBooks and Physical Books Grade 11 Physcs Paper 2
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Grade 11 Physcs Paper 2
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Grade 11 Physcs Paper 2
- Setting Reading Goals Grade 11 Physcs Paper 2
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Grade 11 Physcs Paper 2
- Fact-Checking eBook Content of Grade 11 Physcs Paper 2
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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