

## Mixed Stoichiometry Practice

Name \_\_\_\_\_

Date \_\_\_\_\_ Period \_\_\_\_\_

Write and/or balance the following equations (remember the diatomic elements and to cross-check charges for ionic compounds!!!) Use the mole ratios from the balanced equations to solve the following stoichiometry problems. Use units and labels in all conversions, and round your answer to sig figs.

1. Potassium chlorate decomposes into potassium chloride and oxygen gas.

**Balanced Equation:**

2. How many moles of oxygen are produced when 3.0 moles of potassium chlorate decompose completely?

3. Butane ( $C_4H_{10}$ ) undergoes combustion.

**Balanced Equation:** \_\_\_\_\_  $C_4H_{10}(l)$  + \_\_\_\_\_  $O_2(g)$  → \_\_\_\_\_  $CO_2(g)$  + \_\_\_\_\_  $H_2O(l)$

4. How many grams of  $CO_2$  are produced when 60 g of  $O_2$  are reacted with an excess of butane?

5. Water decomposes into hydrogen gas and oxygen gas by electrolysis.

**Balanced Equation:**

6. How many grams of hydrogen will be produced when 6.0 moles of oxygen are produced?

7. How many grams of water are required to produce 9.00 grams of hydrogen?

# Mixed Stoichiometry Practice Answer Key

**LL Leslie**



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