

# Stoichiometry - Mixed Conversions Answer Key

1. Use the equation below to answer the following questions:



a. How many grams of  $\text{NH}_3$  are made from 31L of  $\text{N}_2$  gas?

$$\frac{31\text{L N}_2}{1} \times \frac{1\text{mole N}_2}{22.4\text{L N}_2} \times \frac{2\text{mole NH}_3}{1\text{mole N}_2} \times \frac{17\text{g NH}_3}{1\text{mole NH}_3} = 47.3\text{g NH}_3$$

b. How many molecules of  $\text{H}_2$  are needed to produce 60g of  $\text{NH}_3$ ?

$$\frac{60\text{g NH}_3}{1} \times \frac{1\text{mole NH}_3}{17\text{g NH}_3} \times \frac{1\text{mole N}_2}{2\text{mole NH}_3} \times \frac{3\text{mole H}_2}{1\text{mole N}_2} \times \frac{6.02 \times 10^{23}\text{ molecules H}_2}{1\text{mole H}_2} = 1.22 \times 10^{24}\text{ molecules H}_2$$

c. How many moles of  $\text{N}_2$  are needed to completely react with 87L of  $\text{H}_2$  gas?

$$\frac{87\text{L}}{1} \times \frac{1\text{mole H}_2}{22.4\text{L H}_2} \times \frac{1\text{mole N}_2}{3\text{mole H}_2} = 1.27\text{mole N}_2$$

2. Butane reacts with oxygen to form water and carbon dioxide.



a. How many molecules of  $\text{H}_2\text{O}$  are produced with 20.7L of  $\text{O}_2$  gas that are completely used during the combustion of butane?

$$\frac{20.7\text{L O}_2}{1} \times \frac{1\text{mole O}_2}{22.4\text{L O}_2} \times \frac{10\text{mole H}_2\text{O}}{13\text{mole O}_2} \times \frac{6.02 \times 10^{23}\text{ molecules H}_2\text{O}}{1\text{mole H}_2\text{O}} = 4.17 \times 10^{23}\text{ molecules H}_2\text{O}$$

b. If 48g of  $\text{H}_2\text{O}$  are collected from the burning of butane, then how many moles of butane were reacted with oxygen?

$$\frac{48\text{g H}_2\text{O}}{1} \times \frac{1\text{mole H}_2\text{O}}{18\text{g H}_2\text{O}} \times \frac{2\text{mole C}_4\text{H}_{10}}{10\text{mole H}_2\text{O}} = 0.53\text{mole C}_4\text{H}_{10}$$

3. Potassium reacts with water to produce potassium hydroxide and hydrogen gas.



a. How many grams of  $\text{H}_2$  will be produced if  $7.22 \times 10^{23}$  atoms of potassium react completely with water?

$$\frac{7.22 \times 10^{23}\text{ atoms K}}{1} \times \frac{1\text{mole K}}{6.02 \times 10^{23}\text{ atoms K}} \times \frac{1\text{mole H}_2}{2\text{mole K}} \times \frac{2\text{g H}_2}{1\text{mole H}_2} = 0.60\text{g H}_2$$

b. How many atoms of K are needed to completely react with  $3.22 \times 10^{23}$  molecules of water?

$$\frac{3.22 \times 10^{23}\text{ molecules H}_2\text{O}}{1} \times \frac{1\text{mole H}_2\text{O}}{6.02 \times 10^{23}\text{ molecules H}_2\text{O}} \times \frac{2\text{mole K}}{2\text{mole H}_2\text{O}} \times \frac{6.02 \times 10^{23}\text{ atoms K}}{1\text{mole K}} = 3.22 \times 10^{23}\text{ atoms K}$$

# Mixed Stoichiometry Worksheet

**CL Gary**



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