

21 This is a question about Group 2 compounds.

Limewater is a solution of calcium hydroxide, commonly used in the identification of carbon dioxide gas. Since calcium hydroxide is only sparingly soluble in water, technicians often make the solution by adding an excess of the solid calcium hydroxide to the required volume of deionised water, shaking the container and then leaving the mixture to settle. In this way, a saturated solution is produced but it can be of variable concentration.

Two students were each given a sample of limewater, from the same batch, in order to determine its concentration. Using 50.0 cm³ portions of the limewater, they carried out titrations using 0.100 mol dm⁻³ hydrochloric acid. One of the students obtained the following results:

Titration	Trial	1	2
Final Volume /cm ³	14.50	28.60	42.70
Initial Volume /cm ³	0.00	14.50	28.60
Volume Added /cm ³	14.50	14.10	14.10

The student decided that the mean titre was 14.10 cm³

The equation for the reaction is:



(a) (i) Calculate the number of moles of hydrochloric acid that reacted.

(1)

(ii) Calculate the number of moles of calcium hydroxide, Ca(OH)₂, that reacted with the acid.

(1)



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**Elizabeth de Freitas, Nathalie
Sinclair, Alf Coles**



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Table of Contents Mmaths Edexcel 2013 Jun

1. Understanding the eBook Mmaths Edexcel 2013 Jun
 - The Rise of Digital Reading Mmaths Edexcel 2013 Jun
 - Advantages of eBooks Over Traditional Books
2. Identifying Mmaths Edexcel 2013 Jun
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mmaths Edexcel 2013 Jun
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mmaths Edexcel 2013 Jun
 - Personalized Recommendations
 - Mmaths Edexcel 2013 Jun User Reviews and Ratings
 - Mmaths Edexcel 2013 Jun and Bestseller Lists

5. Accessing Mmaths Edexcel 2013 Jun Free and Paid eBooks
 - Mmaths Edexcel 2013 Jun Public Domain eBooks
 - Mmaths Edexcel 2013 Jun eBook Subscription Services
 - Mmaths Edexcel 2013 Jun Budget-Friendly Options
6. Navigating Mmaths Edexcel 2013 Jun eBook Formats
 - ePub, PDF, MOBI, and More
 - Mmaths Edexcel 2013 Jun Compatibility with Devices
 - Mmaths Edexcel 2013 Jun Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mmaths Edexcel 2013 Jun
 - Highlighting and Note-Taking Mmaths Edexcel 2013 Jun
 - Interactive Elements Mmaths Edexcel 2013 Jun
8. Staying Engaged with Mmaths Edexcel 2013 Jun
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mmaths Edexcel 2013 Jun
9. Balancing eBooks and Physical Books Mmaths Edexcel 2013 Jun
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mmaths Edexcel 2013 Jun
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Mmaths Edexcel 2013 Jun
 - Setting Reading Goals Mmaths Edexcel 2013 Jun
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Mmaths Edexcel 2013 Jun
 - Fact-Checking eBook Content of Mmaths Edexcel 2013 Jun
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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