

## Chemistry – Unit 5 Worksheet 2

1. An old (pre-1987) penny is nearly pure copper. If such a penny has a mass of 3.3 g, how many moles of copper atoms would be in one penny?

$$3.3 \text{ g} \times \frac{1 \text{ mole}}{63.5 \text{ g}} = 5.2 \times 10^{-2} \text{ mol or } 0.052 \text{ mol}$$

2. Four nails have a total mass of 4.42 grams. How many moles of iron atoms do they contain?

$$4.42 \text{ g} \times \frac{1 \text{ mole}}{55.8 \text{ g}} = 7.92 \times 10^{-2} \text{ mol or } 0.0792 \text{ mol}$$

3. A raindrop has a mass of 0.050 g. How many moles of water does a raindrop contain?

$$0.050 \text{ g} \times \frac{1 \text{ mole}}{18.0 \text{ g}} = 2.78 \times 10^{-3} \text{ mol} \rightarrow 2.8 \times 10^{-3} \text{ mol}$$

4. What mass of water would you need to have 15.0 moles of H<sub>2</sub>O?

$$15.0 \text{ moles} \times \frac{18.0 \text{ g}}{1 \text{ mole}} = 2.70 \times 10^2 \text{ g}$$

5. One box of Morton's Salt contains 737 grams. How many moles of sodium chloride is this?

$$737 \text{ g} \times \frac{1 \text{ mole}}{58.5 \text{ g}} = 12.6 \text{ mol}$$

Na	23.0
Cl	35.5
	<b>58.5 g</b>

6. A chocolate chip cookie recipe calls for 0.050 moles of baking soda (sodium bicarbonate). How many grams should the chef mass out?

$$0.050 \text{ moles} \times \frac{84.0 \text{ g}}{1 \text{ mole}} = 4.2 \text{ g}$$

Na	23.0
H	1.0
C	12.0
O	48.0
	<b>84.0</b>

7. Rust is iron(III) oxide. The owner of a 1959 Cadillac convertible wants to restore it by removing the rust with oxalic acid, but he needs to know how many moles of rust will be involved in the reaction. How many moles of iron(III) oxide are contained in 2.50 kg of rust?

$$2.50 \text{ kg} \times \frac{1000 \text{ g}}{1 \text{ kg}} \times \frac{1 \text{ mole}}{159.6 \text{ g}} = 15.7 \text{ mol Fe}_2\text{O}_3$$

Fe	111.6
O	48.0
	<b>159.6</b>

# Modeling Amta 2013 Answer

**Goyal, Anita**



## **Modeling Amta 2013 Answer:**

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**Table of Contents Modeling Amta 2013 Answer**

1. Understanding the eBook Modeling Amta 2013 Answer
  - The Rise of Digital Reading Modeling Amta 2013 Answer
  - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Amta 2013 Answer
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Modeling Amta 2013 Answer
  - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Amta 2013 Answer
  - Personalized Recommendations
  - Modeling Amta 2013 Answer User Reviews and Ratings
  - Modeling Amta 2013 Answer and Bestseller Lists
5. Accessing Modeling Amta 2013 Answer Free and Paid eBooks
  - Modeling Amta 2013 Answer Public Domain eBooks
  - Modeling Amta 2013 Answer eBook Subscription Services
  - Modeling Amta 2013 Answer Budget-Friendly Options
6. Navigating Modeling Amta 2013 Answer eBook Formats
  - ePub, PDF, MOBI, and More
  - Modeling Amta 2013 Answer Compatibility with Devices
  - Modeling Amta 2013 Answer Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Modeling Amta 2013 Answer
  - Highlighting and Note-Taking Modeling Amta 2013 Answer
  - Interactive Elements Modeling Amta 2013 Answer
8. Staying Engaged with Modeling Amta 2013 Answer

- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Modeling Amta 2013 Answer
9. Balancing eBooks and Physical Books Modeling Amta 2013 Answer
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Modeling Amta 2013 Answer
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Modeling Amta 2013 Answer
- Setting Reading Goals Modeling Amta 2013 Answer
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Modeling Amta 2013 Answer
- Fact-Checking eBook Content of Modeling Amta 2013 Answer
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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