

Name \_\_\_\_\_

Date \_\_\_\_\_ Per \_\_\_\_\_

### Mixed Practice: Names and Formulas

First determine whether each of the following compounds is *ionic* ( $M + NM$ ), *polyatomic* (3 or more atoms), or *covalent* (2  $NM$ 's). Then write the name for each compound.

1)  $\text{NO}_2$  \_\_\_\_\_

2)  $\text{NaBr}$  \_\_\_\_\_

3)  $\text{SiO}_2$  \_\_\_\_\_

4)  $\text{P}_2\text{Br}_4$  \_\_\_\_\_

5)  $\text{FeSO}_4$  \_\_\_\_\_

6)  $\text{SF}_6$  \_\_\_\_\_

7)  $\text{Li}_2\text{S}$  \_\_\_\_\_

8)  $\text{MgBr}_2$  \_\_\_\_\_

9)  $\text{N}_2\text{S}$  \_\_\_\_\_

10)  $\text{Be}(\text{OH})_2$  \_\_\_\_\_

11)  $\text{SO}_3$  \_\_\_\_\_

12)  $\text{Cu}_2\text{S}$  \_\_\_\_\_

13)  $\text{BF}_3$  \_\_\_\_\_

14)  $\text{I}_2$  \_\_\_\_\_

15)  $\text{Pb}_3(\text{PO}_3)_2$  \_\_\_\_\_

# Modeling Chemistry Unit 4 Test

**S Nieto**



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