

Post-lab: Intermolecular Forces

Introduction

The purpose of this experiment is to observe the temperature change of various liquids as they evaporate at different temperatures through a series of beakers. There are a number of intermolecular forces that affect the boiling point of a liquid. The intermolecular forces that affect the boiling point of a liquid are the attractive forces between the molecules of the liquid. In a liquid, the molecules are in constant motion and are held together by attractive forces. The stronger the attractive forces, the higher the boiling point. The weaker the attractive forces, the lower the boiling point. The boiling point of a liquid is the temperature at which the liquid changes from a liquid to a gas. The boiling point of a liquid is affected by the intermolecular forces between the molecules of the liquid. The boiling point of a liquid is also affected by the external pressure. The boiling point of a liquid is the temperature at which the liquid changes from a liquid to a gas. The boiling point of a liquid is affected by the intermolecular forces between the molecules of the liquid. The boiling point of a liquid is also affected by the external pressure. The boiling point of a liquid is the temperature at which the liquid changes from a liquid to a gas.

Materials and Methods (to be in Appendix)

Beaker 1 - 100 mL water and 10 mL of 10% NaCl solution. Beaker 2 - 100 mL water and 10 mL of 10% NaCl solution. Beaker 3 - 100 mL water and 10 mL of 10% NaCl solution. Beaker 4 - 100 mL water and 10 mL of 10% NaCl solution. Beaker 5 - 100 mL water and 10 mL of 10% NaCl solution. Beaker 6 - 100 mL water and 10 mL of 10% NaCl solution. Beaker 7 - 100 mL water and 10 mL of 10% NaCl solution. Beaker 8 - 100 mL water and 10 mL of 10% NaCl solution. Beaker 9 - 100 mL water and 10 mL of 10% NaCl solution. Beaker 10 - 100 mL water and 10 mL of 10% NaCl solution.

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