

CHEM 1202 - Homework # 6

ANSWER KEY

Acids & Bases # 1

Due Tuesday, Nov 9th, 2004 (by Noon)

Check the box to the right if you want your graded homework to be placed out in the public rack outside Prof. Stanley's office. Otherwise you will have to pick up your homework from Prof. Stanley in person:

1. (3 pts) Consider the following weak acids and their pK_a values:



Which compound is the strongest acid (use letter)? **F**

The strongest acid has the smallest or most negative pK_a value. Triflic acid's pK_a of -10 is equal to a $K_a = 1 \times 10^{10}$.

Which compound is the weakest acid (use letter)? **D**

The weakest acid has the largest most positive pK_a value. HOCl acid's pK_a of 7.4 is equal to $K_a = 2.5 \times 10^{-8}$.

Which compound has the strongest conjugate base (use letter)? **D**

The weakest acid has the strongest conjugate base!

2. (5 pts) What are the pH's for the following solutions?

a) 0.001 M HBr = **3**

b) 1 M HNO_3 = **0**

c) 1×10^{-12} M $\text{Sr}(\text{OH})_2$ = **7**

d) 10 M NaOH = **15**

e) 1×10^{-5} M KOH = **9**

See Acids & Bases page 19 in your notes!
 $[\text{OH}^-]_{\text{total}} = [\text{OH}^-]_{\text{acid}} + [\text{OH}^-]_{\text{base}}$
 $[\text{OH}^-]_{\text{total}} = 1 \times 10^{-7}$, but $[\text{OH}^-]_{\text{base}} = 2 \times 10^{-12}$
 which is so small that you can ignore it

3. (3 pts) Calculate the pH of a 0.001 M solution of acid that has a pK_a of 5.0. Clearly show all your work and put a box around your answer.



First, convert the pK_a of 5 into a K_a value: $K_a = \text{antilog}(-pK_a) = 1 \times 10^{-5}$ Now, we can solve for the $[\text{H}^+]$ and the pH.

$$K_a = \frac{(x)(x)}{(0.001 - x)} = 1 \times 10^{-5} \quad \text{assume that } x \ll 0.001, \quad \frac{x^2}{(0.001)} = 1 \times 10^{-5}, \quad \text{or } x^2 = 1 \times 10^{-8}$$

or $x = [\text{H}^+] = 1 \times 10^{-4}$ So the pH = $-\log(1 \times 10^{-4}) = 4$

pH = 4

4. (4 pts) What is the K_a value of a 1 M solution of sabonic acid (HA) that has a pH of 4? Clearly show all your work and put a box around your answer.



BUT, before you go any further, I've given you the pH of the solution in the problem!! So you don't have to solve for x , you already know it!! So all we have to do is take the anti-log of pH 4 to get the concentration of the H^+ (and A^-):

$[\text{H}^+] = \text{antilog}(-4) = 1 \times 10^{-4} \text{ M}$ Substituting this into the equilibrium expression gives us:

$$K_a = \frac{(1 \times 10^{-4})^2}{1 - 1 \times 10^{-4}} = \frac{1 \times 10^{-8}}{1} = 1 \times 10^{-8} \quad \text{so, } K_a = 1 \times 10^{-8}$$

Note that it is fine to drop the x in the $1 - x$ expression because x is much smaller than 1.

So the answer is: $K_a = 1 \times 10^{-8}$

Evan Silberstein Chemistry Answer Key

Tobias Bleicker



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Table of Contents Evan Silberstein Chemistry Answer Key

1. Understanding the eBook Evan Silberstein Chemistry Answer Key
 - The Rise of Digital Reading Evan Silberstein Chemistry Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Evan Silberstein Chemistry Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Evan Silberstein Chemistry Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Evan Silberstein Chemistry Answer Key

- Personalized Recommendations
 - Evan Silberstein Chemistry Answer Key User Reviews and Ratings
 - Evan Silberstein Chemistry Answer Key and Bestseller Lists
5. Accessing Evan Silberstein Chemistry Answer Key Free and Paid eBooks
- Evan Silberstein Chemistry Answer Key Public Domain eBooks
 - Evan Silberstein Chemistry Answer Key eBook Subscription Services
 - Evan Silberstein Chemistry Answer Key Budget-Friendly Options
6. Navigating Evan Silberstein Chemistry Answer Key eBook Formats
- ePub, PDF, MOBI, and More
 - Evan Silberstein Chemistry Answer Key Compatibility with Devices
 - Evan Silberstein Chemistry Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Evan Silberstein Chemistry Answer Key
 - Highlighting and Note-Taking Evan Silberstein Chemistry Answer Key
 - Interactive Elements Evan Silberstein Chemistry Answer Key
8. Staying Engaged with Evan Silberstein Chemistry Answer Key
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Evan Silberstein Chemistry Answer Key
9. Balancing eBooks and Physical Books Evan Silberstein Chemistry Answer Key
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Evan Silberstein Chemistry Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Evan Silberstein Chemistry Answer Key
- Setting Reading Goals Evan Silberstein Chemistry Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Evan Silberstein Chemistry Answer Key

- Fact-Checking eBook Content of Evan Silberstein Chemistry Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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