Fall Semester Review – IPC

Name: _____

Period: _____

Fill in the table at the right from your periodic table.

Element	Chemical Symbol	Atomic #	Valence Electrons	Oxidation #	# of Protons	Atomic Mass	Metal or Nonmetal	Gains or Loses Electrons	# of Full Electron Levels
Oxygen									
Argon									
Calcium									
Nitrogen									
Potassium									
Sulfur									
Helium									

Find the molecular mass of the following compounds

Calcium oxide (CaO)   Water

Classify (ionic, covalent or polyatomic) and name the following compounds.

1. O_2 _____

2. P_2O_5 _____

3. $BrClO_3$ _____

Draw the Lewis Dot Diagrams for:

1. Carbon   2. Sodium

Make Balanced Ionic Compounds from:

1. Li + O _____   3. Be + NO_3^- _____

2. Na + Ar _____   4. Ca + O _____

With dot diagrams draw the covalent bond of O_2:   With dot diagrams the covalent bond of OF_2:

Matching:

A. Mixture
B. Compound
C. Element
D. Matter
E. Heterogeneous
F. Homogeneous

___ A substance made up of two or more atoms that can't be separated by chemical means.

___ Something made up of 2 or more substances that can be separated by physical means.

___ A substance that is pure and made up of only one type of atom.

___ Anything that has mass and takes up space.

___ A mixture that is the same throughout.

___ A mixture that is different throughout.

1. Valence electrons are the outermost electrons of an atom that are involved in chemical bonding? True/False

2. Lithium has _____ valence electrons, will _____ electrons and become (positive or negative). This is why lithium's oxidation number is _____.

3. A metal and non-metal will form a _____ compound; two non-metals form a _____ compound.

Atom A   Atom B
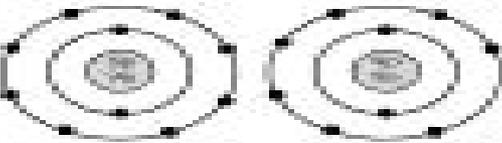
Which one is the ion?

Give the ion notation for the ion.

Matching:

A. Ion
B. Element
C. Isotope
D. Cell

___ If you change the number of protons, you change the _____.

___ If you change the number of neutrons, you change the _____.

___ If you change the number of electrons, you change the _____.

___ Every 3,000 miles you should change a car's _____.

Matching:

A. Law of Conservation of Mass
B. Meter
C. Centimeter
D. Liter
E. Gram

___ Unit of length equal to 3.3 ft.

___ Mass is neither destroyed or created in chemical reaction.

___ Unit of mass; about 1 dollar bill.

___ 1/100th of meter; width of pinky finger.

___ Unit of volume; just bigger than a quart.

Matching:

A. Solid
B. Liquid
C. Gas
D. Oil

___ Molecules that are tightly packed and retain their shape and size.

___ Molecules that bounce off of each other, can be compressed, and take the shape of their container.

___ Molecules that can move (slide) over each other and have a definite size (volume), but not shape and can not be compressed.

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Ying-Ying Zheng

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