

# Final study guide question and answers

What kind of equilibrium exists and under what conditions? **Chemical equilibrium** is a dynamic equilibrium between two or more states of matter, where the concentrations of the reactants and products are constant over time.

What happens if you change the conditions? **Le Chatelier's principle** states that if a system at equilibrium is subjected to a change in concentration, temperature, or pressure, the system will shift in the direction that counteracts the change.

What happens if you change the volume of the container? **Volume change** affects the equilibrium position of a reaction involving gases. If the volume is increased, the equilibrium shifts towards the side with more moles of gas.

What happens if you change the temperature? **Temperature change** affects the equilibrium position of a reaction. For an exothermic reaction, increasing the temperature shifts the equilibrium towards the reactants.

What happens if you change the concentration of the reactants or products? **Concentration change** affects the equilibrium position of a reaction. Increasing the concentration of a reactant shifts the equilibrium towards the products.

What happens if you change the pressure of the system? **Pressure change** affects the equilibrium position of a reaction involving gases. Increasing the pressure shifts the equilibrium towards the side with fewer moles of gas.

What happens if you change the catalyst? **Catalyst** does not affect the equilibrium position of a reaction, but it speeds up the rate at which equilibrium is reached.

What happens if you change the initial concentrations? **Initial concentrations** affect the equilibrium position of a reaction. Increasing the initial concentration of a reactant shifts the equilibrium towards the products.

What happens if you change the initial volume of the container? **Initial volume** affects the equilibrium position of a reaction involving gases. Increasing the initial volume shifts the equilibrium towards the side with more moles of gas.

What happens if you change the initial temperature? **Initial temperature** affects the equilibrium position of a reaction. For an exothermic reaction, increasing the initial temperature shifts the equilibrium towards the reactants.

What happens if you change the initial pressure of the system? **Initial pressure** affects the equilibrium position of a reaction involving gases. Increasing the initial pressure shifts the equilibrium towards the side with fewer moles of gas.

What happens if you change the initial concentration of the reactants or products? **Initial concentration** affects the equilibrium position of a reaction. Increasing the initial concentration of a reactant shifts the equilibrium towards the products.

What happens if you change the initial volume of the container? **Initial volume** affects the equilibrium position of a reaction involving gases. Increasing the initial volume shifts the equilibrium towards the side with more moles of gas.

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What happens if you change the initial pressure of the system? **Initial pressure** affects the equilibrium position of a reaction involving gases. Increasing the initial pressure shifts the equilibrium towards the side with fewer moles of gas.

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web nov 11 2023 2 bring your legs down one at a time when you re ready to exit your handstand coming down is exactly like going up only in reverse extend one leg and carefully lower it down in front of your body as you shift your weight into your palms touch down with the ball of your foot then lower your other leg

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web jun 26 2023 here are the steps start in a lunge your arms should be straight above your head and your dominant leg should be in front slightly bent kick up to handstand you want to feel vertical and try to hold the straight body position by

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