

CHEMICAL KINETICS

EXPERIMENT 24

REPORT SHEET

Name Nicole Frechette

Drawer No. 497

Date 10/2/2020

Partner's Name _____

A. DETERMINATION OF THE ORDER OF THE REACTION

Run #	Time (seconds)	Rate
		a
1	93	0.012
2	181	0.006
3	185	0.005
4	90	0.011

Original Molarity: KI 0.0500M H₂O₂ 0.100M HCl 0.100M

Run #	I ⁻ (KI)	H ₂ O ₂	H ⁺ (HCl)
	Diluted Molarity	Diluted Molarity	Diluted Molarity
1	0.01	0.04	0.02
2	0.005	0.04	0.02
3	0.01	0.02	0.02
4	0.01	0.04	0.01

Calculations:

A. Rate = $\frac{1}{\text{time}(s)}$

Run #

1
2
3

$\frac{1}{93(s)} = 0.012$
 $\frac{1}{181(s)} = 0.006$
 $\frac{1}{185(s)} = 0.005$

Run 1 $\frac{(0.005M)(10mL)}{50} = \frac{M_2(50mL)}{50}$

$(0.005M)(0.2) = M_2 = 0.01 \text{ KI}$

4 $\frac{1}{90(s)} = 0.011$

Kinetics Of A Reaction Lab Report Answers

Kathleen Armour



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