

INTRO TO IONIC BONDING

GUIDED NOTES



Ionic Bonding

All ions have a charge (lost electrons = positive charge, gained electrons = negative charge). How do they form? Describe how they are prepared (electron transfer).

Write down the names of the following elements. Describe how they will behave in reactions.

GROUP 1
Metals (sodium, potassium, lithium, rubidium, caesium, francium). They are soft, shiny, and conduct electricity. They react with water to form hydroxides and hydrogen gas.

GROUP 2
Metals (beryllium, magnesium, calcium, strontium, barium, radium). They are harder than group 1 metals and react with water to form hydroxides and hydrogen gas.

GROUP 7
Non-metals (fluorine, chlorine, bromine, iodine, astatine). They are diatomic molecules and are highly reactive. They react with metals to form ionic compounds.

GROUP 0
Noble gases (helium, neon, argon, krypton, xenon, radon). They are monatomic and are very unreactive.

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Self Assessment Qs

IONIC BONDING

Name: _____

Directions: Answer the following.

Write W for metal and N for non-metal for each of the following.

- _____ gains electrons when bonding
- _____ found on the left side of the periodic table
- _____ has just a few valence electrons
- _____ aluminum
- _____ has more than four valence electrons
- _____ loses electrons when bonding
- _____ found on the right side of the periodic table
- _____ bromine

Draw a Lewis structure for each of the following ionic compounds.

Magnesium + Oxygen	Hydrogen + Chlorine
Lithium + Oxygen	Calcium + Phosphorus

Classwork
Worksheet



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