

# Chemistry: Ionic & Covalent Bonds Unit Study Guide

## Covalent Bonding & Nomenclature Study Guide

1. Circle the correct word that would complete the statements below regarding Covalent Compounds.

- Bond (ion)
- Large (small)
- Atom (molecule)
- High (low)
- Conduct (does not conduct)
- Solid (liquid)
- Soluble (insoluble)

2. Write the prefix following number.

Number	Prefix
1	
2	
3	
4	
5	

3. What are the ions?

4. Complete the chart.

Formula	Name
$Mg^{2+}$	
	$Br^-$
$Fe^{3+}$	
	$Cl^-$

5. What does VSEPR

Name: \_\_\_\_\_ Date: \_\_\_\_\_ Period: \_\_\_\_\_

## Ionic Bonding & Nomenclature Study Guide

1. Circle the correct word that would complete the statements below regarding Ionic Compounds.

- Bond (between a Metal / Nonmetal) and (Metal / Nonmetal)
- Large (small) Electronegativity Difference
- The (Metal / Nonmetal) form Cations which (Gain / Lose) Electrons
- The (Metal / Nonmetal) form Anions which (Gain / Lose) Electrons
- High (Low) Melting & Boiling Point
- Conducts (Does not Conduct) Electricity in Solution
- Solid (Liquid) (Gas) at Room Temperature

2. What charge would the following periodic table families have in a chemical bond?

Family	Charge
Alkali Metals	
Alkali Earths	
Halides	
Carbon	
Nitrogen	
Oxygen	

3. Write the number numbers that is represented by the following numbers.

Number	Prefix
1	
2	
3	
4	
5	
6	

4. Complete the chart below by filling in the missing names.

CATION	ANION	FORMULA	NAME
$Mg^{2+}$	$Br^-$		
		$CuO$	
			Iron (III) sulfate
		$(NH_4)_3PO_4$	

# Ionic Covalent Bonding Test Study Guide Answers

**VM Jensen**



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