

Classifying Chemical Reactions

Analyzing and Predicting Products

Introduction

The power of chemical reactions to transform our lives is visible all around us—in our homes, in our cars, even in our bodies. Chemists try to make sense of the great variety of chemical reactions the same way that biologists organize their knowledge of life, by writing reactions into groups and classifying them. Classifying chemical reactions allows us to predict what chemical reactions will occur when different substances are mixed.

Concepts

- Chemical reactions
- Combination vs. decomposition
- Single vs. double replacement
- Combination reactions

Background

A chemical reaction is defined as any process in which one or more substances are converted into new substances with different properties. Chemical reactions change the identity of the reacting substance(s) and produce new substances. Observing the properties of the reactants and products is therefore a key step in identifying chemical reactions. Some of the observations that may be associated with a chemical reaction include: (1) release of a gas, (2) formation of a precipitate, (3) color changes, (4) temperature changes, and (5) emission or absorption of light. As these observations suggest, chemical reactions can be dynamic and exciting events. The evidence of any chemical reaction—reactants being transformed into products—is summarized in the form of a chemical equation. Consider the reaction represented by Equation 1, the burning of natural gas (methane, CH_4) in a laboratory burner.



The reactants—more specifically, their formulas—are written on the left side of the equation, the products on the right side of the equation. An arrow represents the direction of the reaction and is read as “yields” or “produces.” Other symbols are used to describe the physical state of the reactants and products and the reaction conditions (see Table 1).

Table 1. Symbols in Chemical Equations

| Symbol | Translation |
|----------------------------|--|
| \rightarrow | Yields or produces (separates reactants from products) |
| + | Reacts with or forms alongside (separates two or more reactants or products) |
| Δ | Reaction mixture is heated (written over the arrow) |
| NR | No reaction takes place when reactants are mixed |
| (s) | Pure substance (reactant or product) is a solid |
| (l) | Pure substance (reactant or product) is a liquid |
| (g) | Pure substance (reactant or product) is a gas |
| (aq) | An ionic solution (reactant or product) is dissolved in water |
| $\xrightarrow{\text{Cat}}$ | Catalyst—a substance needed to initiate a reaction. Formula is written over the arrow |

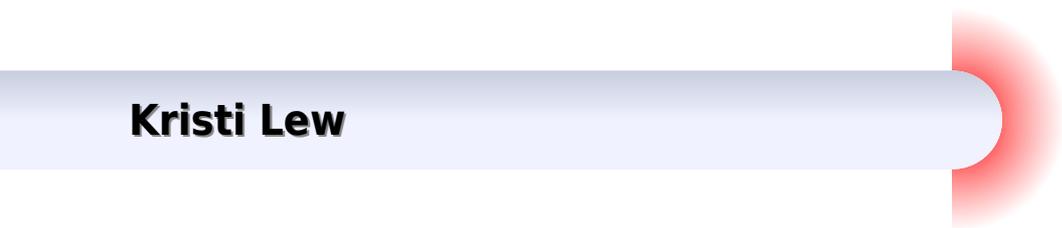
Chemical reactions may be classified by considering the number of reactants and products in the reaction, the physical or chemical nature of the reactants and products, and the rearrangement of atoms in the conversion of the reactants into products (see Table 2 on page 2).

Experiment Overview

The purpose of this experiment is to observe a variety of chemical reactions and to identify patterns in the conversion of reactants into products. The properties of the reactions will be analyzed to classify the chemical reactions into different groups.

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Kristi Lew



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