

# Introduction to Electrochemistry

## Electrolysis of Water

### Introduction

Electrochemistry is the study of the relationship between electrical forces and chemical reactions. There are two basic types of electrochemical processes. In a voltaic cell, commonly known as a battery, the chemical energy from a spontaneous oxidation-reduction reaction is converted into electrical energy. In an electrolytic cell, electricity from an external source is used to “force” a nonspontaneous chemical reaction to occur. What chemical reaction will take place when an electric current flows through water?

### Concepts

- Electrochemistry
- Electrolysis
- Oxidation-reduction
- Anodes and cathodes

### Background

The first electrochemical process to produce electricity was described in 1800 by the Italian scientist Alessandro Volta, a former high school teacher. Acting on the hypothesis that two dissimilar metals could serve as a source of electricity, Volta constructed a stacked pile of alternating silver and zinc plates separated by pads of absorbent material soaked in saltwater. When Volta moistened his fingers and repeatedly touched the top and bottom metal plates, he experienced a series of small electric shocks. The “voltaic pile,” as it came to be called, was the first battery—a chemical method of generating an electric current. Within months, William Nicholson and Anthony Carlisle in England attempted to confirm the production of electric charges on the upper and lower plates in a voltaic pile using an electroscope. In order to connect the plates to the electroscope, Nicholson and Carlisle added zinc water to the uppermost metal plate and inserted a wire to the electroscope. To their surprise, Nicholson and Carlisle observed the formation of a gas, which they identified as hydrogen. Nicholson and Carlisle then filled a small tube with river water and inserted wires from the voltaic pile into each end of the tube. Two different gases were generated, one at each wire—Nicholson and Carlisle had discovered electrolysis.

### Experiment Overview

The purpose of this experiment is to investigate the electrolysis of water in an electrochemical cell. Two carbon pencil “leads” will be inserted into the opposite ends of a U-tube containing water, sodium sulfate and bromothymol blue. An electric current will be passed through the solution by connecting the pencil leads to the positive and negative terminals of a 9-volt battery (see Figure 1). The pencil leads act as external conductors and provide a surface for the chemical reaction. Sodium sulfate, an ionic compound, is needed to improve the current flow through the solution. Bromothymol blue, an acid-base indicator, will help to identify the changes occurring in the solution as the electrolysis proceeds. Bromothymol blue is yellow in acidic solutions (pH < 6.0), blue in basic solutions (pH > 7.6), and various shades of green at intermediate pH values (pH = 6.0–7.6).

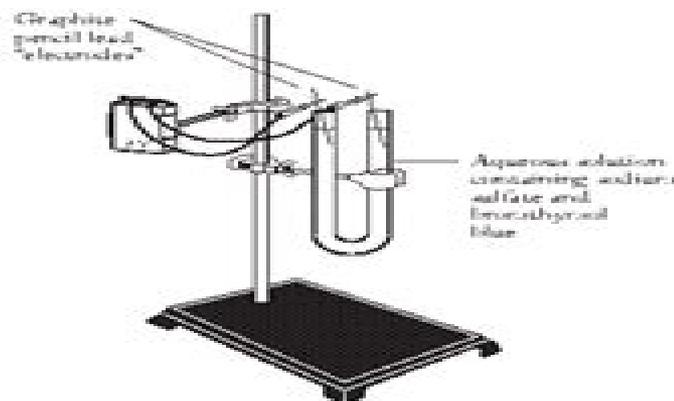


Figure 1. U-tube setup for the electrolysis of water.

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