

Classic Titration

pH Curves and an Unknown

Introduction

One of the most common questions chemists have to answer is how much of something is present in a sample or a product. If the product contains an acid or base, these questions are usually answered by titration. Acid–base titrations can be used to measure the concentration of an acid or base in solution and to calculate the formula (molar) mass of an unknown acid or base.

Concepts

- Acids and bases
- Neutralization reactions
- Titration curve
- Titration
- Equivalence point

Background

Titration is a method of *colorimetric analysis*—the use of volume measurements to analyze an unknown. In acid–base chemistry, titration is most often used to analyze the amount of acid or base in a sample or solution. Consider a solution containing an unknown amount of hydrochloric acid. In a titration experiment, a known volume of the hydrochloric acid solution would be “titrated” by slowly adding (dropwise) a standard solution of a strong base, such as sodium hydroxide. (A standard solution is one with a concentration that is accurately known.) The titrant, sodium hydroxide (in this case), reacts with and consumes the acid via a neutralization reaction (Equation 1). The exact volume of base needed to react completely with the acid is measured. This is called the *equivalence point* of the titration—the point at which stoichiometric amounts of the acid and base have combined.



Knowing the exact concentration and volume added of the titrant gives the number of moles of sodium hydroxide. The latter, in turn, is related by stoichiometry to the number of moles of hydrochloric acid initially present in the unknown.

Indicators are usually added to acid–base titrations to detect the equivalence point. The end point of the titration is the point at which the indicator changes color and signals that the equivalence point has indeed been reached. For example, in the case of the neutralization reaction shown in Reaction 1, the pH of the solution would be acidic (< 7) before the equivalence point and basic (> 7) after the equivalence point. The pH at the equivalence point should be exactly 7, corresponding to the neutral products (sodium chloride and water). An indicator that changes color around pH 7 is therefore a suitable indicator for the titration of a strong acid with a strong base.

The progress of an acid–base titration can also be followed by measuring the pH of the solution being analyzed as a function of the volume of titrant added. A plot of the resulting data is called a pH curve or titration curve. Titration curves allow a precise determination of the equivalence point of the titration without the use of an indicator.

Experiment Overview

The purpose of this experiment is to analyze the shape of the titration curve for neutralization of hydrochloric acid with sodium hydroxide and to determine the concentration of an unknown hydrochloric acid solution. In Part A, pH measurements may be made electronically as a function of time using a pH sensor attached to a computer- or calculator-based interface system, or manually as a function of volume added using a pH meter. In Part B, the titration will be repeated with the use of an indicator to assess the endpoint.

Pre-Lab Questions

The approximate concentration of a hydrochloric acid solution is 0.5 M. The exact concentration of this solution is to be determined by titration with 0.215 M sodium hydroxide solution.

1. A 10.00-mL sample of the HCl solution was transferred by pipet to an Erlenmeyer flask and then diluted by adding about 40 mL of distilled water. What is the approximate H_3O^+ concentration and pH of the solution in the flask before the titration begins?
2. Thymolphthalein indicator was added, and the solution in the flask was titrated with 0.215 M NaOH until the indicator just turned pink (pH = 8–9). The exact volume of NaOH required was 22.75 mL. Use the following equation to calculate the concentration of HCl in the original 10.00-mL sample.

$$M_1 \times V_1 = nM_2 \times V_2$$

- M_1 = molarity of standard base solution
- M_2 = unknown molarity of acid solution
- V_1 = volume of base added
- V_2 = initial volume of acid solution
- n = mole ratio (number of moles of base that react with one mole of acid)

3. One student accidentally “overshot” the endpoint and added 23.50 mL of 0.215 M NaOH. Is the calculated concentration of HCl likely to be too high or too low as a result of this error?

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Table of Contents Flinn Target Titration

1. Understanding the eBook Flinn Target Titration
 - The Rise of Digital Reading Flinn Target Titration
 - Advantages of eBooks Over Traditional Books
2. Identifying Flinn Target Titration
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Flinn Target Titration
 - User-Friendly Interface
4. Exploring eBook Recommendations from Flinn Target Titration
 - Personalized Recommendations
 - Flinn Target Titration User Reviews and Ratings
 - Flinn Target Titration and Bestseller Lists

5. Accessing Flinn Target Titration Free and Paid eBooks
 - Flinn Target Titration Public Domain eBooks
 - Flinn Target Titration eBook Subscription Services
 - Flinn Target Titration Budget-Friendly Options
6. Navigating Flinn Target Titration eBook Formats
 - ePub, PDF, MOBI, and More
 - Flinn Target Titration Compatibility with Devices
 - Flinn Target Titration Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Flinn Target Titration
 - Highlighting and Note-Taking Flinn Target Titration
 - Interactive Elements Flinn Target Titration
8. Staying Engaged with Flinn Target Titration
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Flinn Target Titration
9. Balancing eBooks and Physical Books Flinn Target Titration
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Flinn Target Titration
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Flinn Target Titration
 - Setting Reading Goals Flinn Target Titration
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Flinn Target Titration
 - Fact-Checking eBook Content of Flinn Target Titration
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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