

Gas Laws Problem Packet

Gas Laws Packet

Name _____ Date _____ Period _____

Essentials

Standard Constant:

Unit	Standard Pressure
atm	1.013
mmHg	760
cm Hg	76
torr	760
atm	1.013

22.4 L = 1 mol at STP

STP (Standard Temperature and Pressure) = 0°C and 1 atm

The gas laws:

Boyle's Law:	$P_1V_1 = P_2V_2$
Charles's Law:	$\frac{V_1}{T_1} = \frac{V_2}{T_2}$
Gay-Lussac's Law:	$\frac{P_1}{T_1} = \frac{P_2}{T_2}$
Avogadro's Law:	$\frac{V_1}{n_1} = \frac{V_2}{n_2}$
Combined Gas Law:	$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$
Graham's Law:	$\frac{r_1}{r_2} = \sqrt{\frac{M_2}{M_1}}$

The constant R (used in the other gas laws) depends on the units of Pressure:

- 1 atm, R = 0.0821 $\frac{\text{L}\cdot\text{atm}}{\text{mol}\cdot\text{K}}$
- 1 atm, R = 0.0821 $\frac{\text{L}\cdot\text{atm}}{\text{mol}\cdot\text{K}}$

Always Remember:

- Temperature must always be in Kelvin! ($K = ^\circ C + 273$)
- If you are given two different units of pressure, CONVERT (and)
- Create two columns before solving, initial and final, write what you're given in the problem. Find the law that uses the variables you're given, and solve for the unknown. (for Ideal Gas Law, write what you're given, insert those in equation and solve for unknown)
- Avogadro's Law: DON'T FORGET TO USE MOLE (MOL) (shown)
- Graham's Law: You will ALWAYS have the molecular mass for these problems. Double and triple check you have the mole of effluent and molecular mass in the correct spots before solving the equation (this is an INVERSE relationship). You MUST calculate the square root of double both sides of the equation before moving things.

Charles' Law $V \propto T$

Directions: Complete each of the problems below. Show how you set up the problem and solve.

- The temperature inside my refrigerator is about 4°C. If I place a balloon in the refrigerator, will the temperature of the air inside the balloon increase or decrease? Explain your answer. (15 pts)
- A balloon is inflated to a volume of 2.0 L at 25°C. If the temperature is lowered to 10°C, what will the volume of the balloon be? (15 pts)

Gay-Lussac's Law $P \propto T$

- A gas sample is held at a constant volume of 2.0 L and a pressure of 1.0 atm. If the temperature is increased from 25°C to 50°C, what will the pressure be? (15 pts)
- A gas sample is held at a constant volume of 2.0 L and a pressure of 1.0 atm. If the temperature is decreased from 25°C to 10°C, what will the pressure be? (15 pts)

Avogadro's Law $V \propto n$

- A gas sample is held at a constant pressure and temperature of 25°C. If the volume is increased from 2.0 L to 4.0 L, what will the number of moles of gas be? (15 pts)
- A gas sample is held at a constant pressure and temperature of 25°C. If the volume is decreased from 2.0 L to 1.0 L, what will the number of moles of gas be? (15 pts)

Combined Gas Law $\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$

- A gas sample is held at a constant pressure and temperature of 25°C. If the volume is increased from 2.0 L to 4.0 L, what will the number of moles of gas be? (15 pts)
- A gas sample is held at a constant pressure and temperature of 25°C. If the volume is decreased from 2.0 L to 1.0 L, what will the number of moles of gas be? (15 pts)

Graham's Law $\frac{r_1}{r_2} = \sqrt{\frac{M_2}{M_1}}$

- Gas A has a molecular mass of 16 g/mol and a rate of effusion of 1.0 L/min. Gas B has a molecular mass of 32 g/mol. What is the rate of effusion of Gas B? (15 pts)
- Gas A has a molecular mass of 16 g/mol and a rate of effusion of 1.0 L/min. Gas B has a molecular mass of 32 g/mol. What is the rate of effusion of Gas B? (15 pts)

All Gas Laws Included!



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