

Name: _____

Period: _____



- Solute **A**
 - Solvent **B**
 - Covalent **C**
 - Ionic **D**
 - Concentration **E**
- M or N

- The part that dissolves into the solution.
- The part that dissolves something.
- Compounds made up of two non-metals.
- Since they are made up of ions, they are polar molecules.
- Tells you how much solvent is dissolved in the solution.

- Dissolve **E**
- Pressure **A**
- Crushing **C**
- Heating **D**
- Stirring **B**

- Causes more gaseous solvent to be forced into a liquid.
- Increases solubility by moving the molecules around faster.
- Increases solubility by allowing the solvent to touch more sides of the solute.
- Expands the solvent a bit to allow more room for the solute.
- When something seems to disappear inside a solvent.

- Water tension **F**
- Dissociation **A**
- Hydrogen bonds **D**
- Unpolar Molecule **C**
- Capillary action **F**
- Cohesion **C**

- The breaking up of an ionic molecule by a polar solvent.
- A molecule that has a positive and negative end.
- The attractive force between two molecules of the same substance.
- The attractive force caused by hydrogen protons to a more negative non-metal.
- What allows plants to draw water up from their roots.
- Caused by the cohesion of water molecules holding it together.

- Water or Non-polar?
- | | |
|------------------------------|-----------------------------|
| N Cooking oil | P Salt |
| P Soluble in water | N Ionic Compounds |
| P A compound with H's | N Dissolves in water |
| P Sugar | N Insoluble in water |

- | | |
|-----------------------------|-----------------------------|
| P Salt | N Ionic Compounds |
| N Dissolves in water | N Insoluble in water |

Draw a water molecule, being sure to label the ends as positive and negative.



Give three ways to make sugar dissolve faster.

- stirring
- heating
- crushing

Do metals become positive or negative?

Would a metal be attracted to water's hydrogen or oxygen?

Would a nonmetal be attracted to water's hydrogen or oxygen?

Which is the smallest particle size: rock salt, regular salt, or only powdered salt?

Which is the largest particle size: sugar cube, granulated sugar (regular table sugar), or powdered sugar?

To which side of a Water Molecule are these Attracted?

O Magnesium	H Sulfur	O I ₂
H Chlorine	O Potassium	H Bromine
O Calcium	neither Helium	H Nitrogen

Which one will dissolve the fastest?

powdered sugar or granulated sugar?

hot water or cold water?

stirred or not stirred?

Large particles or small particles?

 OR 

Which is warmer: arctic water or tropical water?

Which holds more oxygen: arctic water or tropical water?

When power plants use cold streams for cooling, they can kill the fish by heating up the water. Why do many fish die? lack of oxygen (warm water causes the O₂ to be released in air)

Which holds more solid solute: hot or cold liquids?

Which holds more gaseous solute: hot or cold liquids?

Why are water bugs able to "walk on water"?

water tension (cohesion)

How do plants get water from their roots up to their leaves?

capillary ~~action~~ action

What property of gases allows them to be pressured into solution?

can be compressed

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