

Practical Problems Chapter 9: Solutions

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2. The following procedure illustrates how to determine the molar mass of a volatile liquid. A student used the following procedure to determine the molar mass of a substance with a density of 1.10 g/mL. What is the molar mass of the substance? (M_w = 100 g/mol)

10. procedure

Initial mass of flask: $\left(\frac{100 \text{ g/mol}}{1 \text{ mol/mol}} \right) = 100 \text{ g/mol}$

10. $\left(\frac{100}{100} \right) = 10 \left(\frac{100}{100} \right) = \frac{10000}{10000} \text{ mol/L} \times \frac{1000 \text{ mL}}{1000 \text{ mL}} = \frac{10000 \text{ mol}}{10000 \text{ mL}} = 1.00 \text{ mol/L} \left(\frac{100}{100} \right) \text{ mol/L}$

10. procedure

10. $\left(\frac{100}{100} \right) = 10 \left(\frac{100}{100} \right) = \frac{10000}{10000} \text{ mol/L} \times \frac{1000 \text{ mL}}{1000 \text{ mL}} = \frac{10000 \text{ mol}}{10000 \text{ mL}} = 1.00 \text{ mol/L} \left(\frac{100}{100} \right) \text{ mol/L}$

procedure

Initial mass of flask: $\left(\frac{10000 \text{ mol}}{10000 \text{ mL}} \right) = 1.00 \text{ mol/L}$

Initial mass of flask: $\left(\frac{100 \text{ g/mol}}{1 \text{ mol/mol}} \right) = 100 \text{ g/mol}$

Procedure 10: $\frac{100 \text{ g/mol}}{10000 \text{ mL}} = \frac{10000 \text{ mol}}{10000 \text{ mL}} = 1.00 \text{ mol/L}$

10. procedure: $\frac{100 \text{ g/mol}}{10000 \text{ mL}}$

10. procedure: $\frac{10000 \text{ mol}}{10000 \text{ mL}}$

procedure: $\frac{10000 \text{ mol}}{10000 \text{ mL}}$

2. How many grams of NaOH are represented under the flask if a 100 mL solution of NaOH ?

Initial mass of flask: $\left(\frac{100 \text{ g/mol}}{1 \text{ mol/mol}} \right) = 100 \text{ g/mol}$

Mass of: 100 g NaOH

Guided Practice Problems Chapter 8 Chemistry

**Patrick E. McMahon, Rosemary
McMahon, Bohdan Khomtchouk**



Guided Practice Problems Chapter 8 Chemistry:

Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics. The focus is not to recap or review the theoretical concepts well described in the available texts. Instead, the topics and descriptions in this book make available specific, detailed, step-by-step methods and procedures for solving the major types of problems in general chemistry. Explanations, instructional process sequences, solved examples, and completely solved practice problems are greatly expanded, containing significantly more detail than can usually be devoted to in a comprehensive text. Many chapters also provide alternative viewpoints as an aid to understanding. Key Features: The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester. Each is written in a specific and detailed step-by-step process for problem solving, whether mathematical or conceptual. Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts. Includes a chapter designed to eliminate confusion concerning acid-base reactions, which often persists through working with acid-base equilibrium. Many chapters provide alternative viewpoints as an aid to understanding. This book addresses a very real need for a large number of incoming freshmen in STEM fields.

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Merrill Chemistry Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998

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