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CHAPTER 2

ATOMS, MOLECULES, AND IONS

Problem Categories

Biological: 2.79, 2.80.

Conceptual: 2.31, 2.32, 2.33, 2.34, 2.61, 2.65, 2.71, 2.72, 2.81, 2.82, 2.93, 2.103, 2.113.

Descriptive: 2.24, 2.25, 2.26, 2.49, 2.50, 2.66, 2.70, 2.76, 2.84, 2.85, 2.86, 2.87, 2.88, 2.90, 2.91, 2.92, 2.94, 2.98, 2.101, 2.114.

Environmental: 2.122.

Organic: 2.47, 2.48, 2.69, 2.105, 2.107, 2.108, 2.115.

Difficulty Level

Easy: 2.7, 2.8, 2.13, 2.14, 2.15, 2.16, 2.23, 2.31, 2.32, 2.33, 2.43, 2.44, 2.45, 2.46, 2.47, 2.48, 2.62, 2.91, 2.92, 2.99, 2.100.

Medium: 2.17, 2.18, 2.24, 2.26, 2.34, 2.35, 2.36, 2.49, 2.50, 2.57, 2.58, 2.59, 2.60, 2.61, 2.63, 2.64, 2.65, 2.66, 2.67, 2.68, 2.69, 2.70, 2.71, 2.72, 2.73, 2.74, 2.75, 2.76, 2.79, 2.80, 2.81, 2.82, 2.83, 2.84, 2.85, 2.86, 2.88, 2.89, 2.90, 2.93, 2.94, 2.95, 2.96, 2.98, 2.101, 2.102, 2.103, 2.110, 2.112, 2.114, 2.115.

Difficult: 2.25, 2.77, 2.78, 2.87, 2.97, 2.104, 2.105, 2.106, 2.107, 2.108, 2.109, 2.111, 2.113.

First, convert 1 cm to picometers:

$$? \text{ He atoms} = (1 \times 10^{10} \cancel{\text{ pm}}) \times \frac{1 \text{ He atom}}{1 \times 10^2 \cancel{\text{ pm}}} = 1 \times 10^8 \text{ He atoms}$$

Note that you are given information to set up the unit factor relating meters and miles.

$$r_{\text{atom}} = 10^4 r_{\text{nucleus}} = 10^4 \times 2.0 \cancel{\text{ cm}} \times \frac{10^2 \cancel{\text{ m}}}{100 \cancel{\text{ cm}}} \times \frac{1 \text{ mi}}{1609 \cancel{\text{ m}}} = 0.12 \text{ mi}$$

For iron, the atomic number Z is 26. Therefore the mass number A is:

$$A = 26 + 28 = 54$$

Strategy: The 239 in Pu-239 is the mass number. The mass number (A) is the total number of

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Michael Brown



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Student Solution Manual to Accompany Understanding Intermediate Algebra by Lewis Hirsch and Arthur Goodman

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Wagenknecht,Nancy J. Nersessian,Hanne Andersen,2015-06-12 The book examines the emerging approach of using qualitative methods such as interviews and field observations in the philosophy of science Qualitative methods are gaining popularity among philosophers of science as more and more scholars are resorting to empirical work in their study of scientific practices At the same time the results produced through empirical work are quite different from those gained through the kind of introspective conceptual analysis more typical of philosophy This volume explores the benefits and challenges of an empirical philosophy of science and addresses questions such as What do philosophers gain from empirical work How can empirical research help to develop philosophical concepts How do we integrate philosophical frameworks and empirical research What constraints do we accept when choosing an empirical approach What constraints does a pronounced theoretical focus impose on empirical work Nine experts discuss their thoughts and empirical results in the chapters of this book with the aim of providing readers with an answer to these questions *The Nebraska Teacher* ,1916 **The Lincoln**

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Molecular Quantum Mechanics Peter W. Atkins,Ronald S. Friedman,2011 This text unravels those fundamental physical principles which explain how all matter behaves It takes us from the foundations of quantum mechanics through quantum models of atomic molecular and electronic structure and on to discussions of spectroscopy and the electronic and magnetic properties of molecules

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