

General Chemistry II Jasperse  
Acid-Base Chemistry: Extra Practice Problems

General Types/Groups of problems:

Conceptual Questions: Acids, Bases, and Conjugates, Miscellaneous	p1	$K_b$ and $pK_b$ , Base Strength, and using $K_b$ or $pK_b$ to Calculate $[OH^-]$ , $pOH$ , $pH$ , and/or $[H^+]$	p7-10
Recognizing Strong versus Weak Acids; Recognizing Basic versus Nonbasic	p3	Recognizing Acid/Base Properties when Ionics are Dissolved in Water	p11
pH Calculations: Relationships between pH and pOH	p4	Answers	p12
$K_a$ , Sense + Calculations: Using $K_a$ or $pK_a$ to Calculate $[H^+]$ and/or pH; using pH to calculate $K_a$ or $pK_a$	p5		

**Conceptual Questions: Acids, Bases, and Conjugates, Miscellaneous**

- In the Brønsted-Lowry definition of acids and bases, an acid \_\_\_\_\_
  - is a proton donor.
  - is a proton acceptor.
  - forms stable hydrogen bonds.
  - breaks stable hydrogen bonds.
  - corrodes metals.
- In the Brønsted-Lowry definition of acids and bases, a base \_\_\_\_\_
  - is a proton donor.
  - is a proton acceptor.
  - forms stable hydrogen bonds.
  - breaks stable hydrogen bonds.
  - corrodes metals.
- In the following reaction in aqueous solution, the acid reactant is \_\_\_\_\_ and its conjugate base product is \_\_\_\_\_
 
$$CH_3COOH + NH_3 \rightleftharpoons CH_3COO^- + NH_4^+$$
  - $CH_3COOH$ ;  $CH_3COO^-$
  - $CH_3COOH$ ;  $NH_4^+$
  - $NH_3$ ;  $CH_3COO^-$
  - $NH_3$ ;  $NH_4^+$
  - $CH_3COOH$ ;  $H_3O^+$
- In the following reaction in aqueous solution, the acid reactant is \_\_\_\_\_, and its conjugate base product is \_\_\_\_\_
 
$$CH_3NH_2 + HSO_4^- \rightleftharpoons CH_3NH_3^+ + SO_4^{2-}$$
  - $CH_3NH_2$ ;  $CH_3NH_3^+$
  - $CH_3NH_2$ ;  $SO_4^{2-}$
  - $HSO_4^-$ ;  $CH_3NH_3^+$
  - $HSO_4^-$ ;  $SO_4^{2-}$
  - $HSO_4^-$ ;  $H_3O^+$
- Which of the following is the conjugate acid of the hydrogen phosphate ion,  $HPO_4^{2-}$ ?
  - $H_2PO_4^-$
  - $H_3PO_4$
  - $HPO_4^{2-}$
  - $PO_4^{3-}$
  - $H_3O^+$

# General Chemistry Acid Base Practice Problems

**K Morrison**



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