

GENERAL CHEMISTRY
UNIT 4 PART 1: REACTIONS

4.4 TYPES OF CHEMICAL REACTIONS (OVERVIEW)

A. NON-REDOX REACTIONS

- In these reactions, ions and atoms have already achieved stable electron configurations
- Therefore, there are no transfer of e-s in these rxns
- Rearrangement of atoms occur to produce the most stable compounds

B. OXIDATION-REDUCTION (REDOX) RXNS

- In these rxns, ions and atoms have NOT yet achieved stable electron configurations
- Therefore, e-s are transferred in these rxns to achieve the most stable e-configurations

4.6 DESCRIBING CHEMICAL RXNS IN A SOLUTIONS

A. MOLECULAR EQUATION

- Everything is written as a compound even though it may exist as ions in solution
- Ex: $\text{NiCl}_2(\text{aq}) + \text{Zn}(\text{s}) \rightarrow \text{ZnCl}_2(\text{aq}) + \text{Ni}(\text{s})$

B. COMPLETE IONIC EQUATION (CIE)

- The strong electrolytes are written as ions
- Ex: $\text{Ni}^{2+}(\text{aq}) + 2\text{Cl}^{-}(\text{aq}) + \text{Zn}(\text{s}) \rightarrow \text{Zn}^{2+}(\text{aq}) + 2\text{Cl}^{-}(\text{aq})$

C. NET IONIC EQUATION (NIE)

- The ions that don't participate in the rxn (spectator ions) are not shown
- Ex: $\text{Ni}^{2+}(\text{aq}) + \text{Zn}(\text{s}) \rightarrow \text{Zn}^{2+}(\text{aq}) + \text{Ni}(\text{s})$

4.9 OXIDATION REDUCTION RXNS (REDOX)

A. OXIDATION (OX)

- Loss of e-s, substance becomes more positive (Zn is oxidized)
- Ox $\frac{1}{2}$ reaction: shows just what is losing e-s
- Ex: $\text{Zn}(\text{s}) \rightarrow \text{Zn}^{2+}(\text{aq}) + 2\text{e}^{-}$

B. REDUCTION (RED)

- Gain of e-s, substance becomes less positive (Ni^{2+} is reduced)
- Red $\frac{1}{2}$ rxn: shows just what is gaining e-s
- Ex: $\text{Ni}^{2+}(\text{aq}) + 2\text{e}^{-} \rightarrow \text{Ni}(\text{s})$
- OIL RIG (oxidation is losing e-, reduction is gaining e-)

C. OXIDIZING AGENT

- The substance that causes the ox, takes the e-s, is reduced
- Ex: OA = Ni^{2+}

D. REDUCING AGENT

- The substance that causes the red, gives the e-s, is oxidized
- Ex: RA = Zn

E. TYPES OF REDOX RXNS

- Combination/composition/synthesis

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