

1. Consider two acids HCO_2H (formic acid, $\text{Pka} = 3.8$) and pivalic acid $[(\text{CH}_3)_3\text{CCO}_2\text{H}]$, $\text{Pka} = 5.0$ (4)

a. Which acid has the larger K_a ?

pivalic acid

$\text{Pka} = 5.0$

$$\begin{aligned} K_a &= 10^{-\text{pKa}} \\ &= 10^{-5.0} \\ &= 10^{-5} \end{aligned}$$

(0.5)

formic acid

$\text{pka} = 3.8$

$$\begin{aligned} K_a &= 10^{-3.8} \\ &= 1.5 \times 10^{-4} \end{aligned}$$

Formic acid has higher K_a (0.5)

b. Which acid is strong acid?

Formic acid is strong (less Pka) (1)

c. Which acid forms the stronger conjugate base?

Pivalic acid will form a strong conjugate base (1)

d. When each acid is dissolved in water, for which acid does the equilibrium lie further to the right? Illustrate your reasoning with a reaction scheme.



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Christian Drosten



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