

# GRAPHING TRENDS OF THE PERIODIC TABLE

## CHEMISTRY

NAME: \_\_\_\_\_

DATE: \_\_\_\_\_

### PERIOD: \_\_\_\_\_ Graphing Periodic Table Trends

**Directions:** Using the provided graph paper, graph each of the elemental properties using a line graph. Place the property on the Y-axis and the Atomic number on the X-axis. Draw a dark, solid, vertical line on each graph where each row on the periodic table ends. (i.e. after atomic number 20). Use the graphs to analyze the trends in these properties and to answer the questions on the back.

Atomic Radius (pm)							
Lithium - 3	157	Sodium - 11	190	Potassium - 19	243	Rubidium - 27	295
Beryllium - 4	112	Magnesium - 12	145	Calcium - 20	194	Strontium - 28	219
Boron - 5	87	Aluminum - 13	118	Gallium - 31	135	Indium - 40	156
Carbon - 6	67	Silicon - 14	111	Germanium - 32	125	Tin - 50	145
Nitrogen - 7	55	Phosphorus - 15	98	Arsenic - 33	114	Antimony - 51	133
Oxygen - 8	48	Sulfur - 16	88	Selenium - 34	103	Tellurium - 52	123
Fluorine - 9	42	Chlorine - 17	79	Bromine - 35	94	Iodine - 53	119

Ionization Energy (kJ/mol)							
Lithium - 3	520	Sodium - 11	496	Potassium - 19	419	Rubidium - 27	403
Beryllium - 4	900	Magnesium - 12	738	Calcium - 20	590	Strontium - 28	550

### **Atomic Radius**



# Graphing Trends In The Periodic Table Key

**Robert C. Smoot, Smoot, Richard G.  
Smith, Jack Price**



## **Graphing Trends In The Periodic Table Key:**

**Chemistry in the Community (Enhanced Core Four)** American Chemical Society,2006-02-15 *Chemistry Homework for OCR A for Double and Separate Awards* Gareth Pritchard,2001 Howework activites for OCR A Chemistry specifications *ChemDiscovery Student Guide* Olga I. Agapova,2002 *Chemistry for AQA*. Ann Fullick,Patrick Fullick,2001 This resource has separate books for biology chemistry and physics Each book is accompanied by a teacher s resource pack on customizable CD ROM or as a printed pack The series is designed to work in conjunction with the Coordinated Science for AQA series so that coordinated and separate science can be taught alongside each other

**Prentice Hall Science Explorer** Michael J. Padilla,2002 **Catalog of Instructional Films (16mm, Super 8, and 8mm) for College Chemistry** Advisory Council on College Chemistry,1969 *Chemistry in the Community (ChemCom)* American Chemical Society,2011-06-17 Touted as the most successful NSF funded project published Chemistry in the Community ChemCom by the American Chemical Society ACS offers a meaningful and memorable chemistry program for all levels of high school students ChemCom covers traditional chemistry topics within the context of societal issues and real world scenarios Centered on decision making activities where students are responsible for generating data in an investigating analyzing that data and then applying their chemistry knowledge to solve the presented problem The text is intensively laboratory based with all 39 of the investigations integrated within the text not separate from the reading With the ChemCom program students learn more organic and biochemistry more environmental and industrial chemistry and more on the particulate nature of matter than other textbooks all within the relevance of solving problems that arise in everyday life Meticulously updated to meet the needs of today s teachers and students the new sixth edition of ChemCom adheres to the new science framework as well as the forthcoming next generation of science standards Incorporating advances in learning and cognitive sciences ChemCom s wide ranging coverage builds upon the concepts and principles found in the National Science Education Standards Correlations are available showing how closely aligned ChemCom is to these and other state standards *Chemical Structure and Reactivity* James Keeler,Peter Wothers,2013-11 Why do certain substances react together in the way that they do What determines the shape of molecules And how can we predict whether a particular reaction will happen at all Such questions lie at the heart of chemistry the science of understanding the composition of substances their reactions and properties Though introductory chemistry is often broken into three sections inorganic organic and physical the only way for students to fully understand the subject is to see it as a single unified whole Chemical Structure and Reactivity rises to the challenge of depicting the reality of chemistry Offering a fresh approach to the subject by depicting it as a seamless discipline the text shows how organic inorganic and physical concepts can be blended together in order to achieve the common goal of understanding chemical systems With a lively and engaging writing style enhanced by vivid illustrations only Chemical Structure and Reactivity makes teaching chemistry with an integrated

approach possible Special Features The only introductory text to take a truly integrated approach in explaining the fundamentals of chemistry Fosters an orbital based understanding of reactions with clear curly arrow mechanistic detail throughout A two part structure allows flexibility of use Part I lays down the core of the subject while Part II describes a series of relatively standalone topics which can be selected to fit a particular course Numerous concepts are illustrated with fully cross referenced custom developed online modules enabling students to develop an understanding through active learning Self test exercises embedded in the text with solutions at the end of each chapter and extensive question sets encourage hands on learning to help students master the subject and gain confidence The Online Resource Centre features a range of additional resources for both students and registered adopters of the book New to this Edition A new chapter on symmetry has been added to Part I Discussions of organometallic chemistry spectroscopy and molecular geometry have been expanded Cross references from Part I to Part II have been increased to make the links between core concepts and more advanced topics clearer More self test questions and exercises have been provided **Chemistry** Trace Jordan, Neville R. Kallenbach, 2017 *Chemistry The Molecules of Life* offers chemical insights within the context of health pharmaceuticals and the function of biological molecules The contextualized presentation of topics gives students a broad introduction to chemistry and helps them to see the relevance of chemistry to their personal lives **Prentice Hall Chemistry**, 2000

*Holt Chemistry* Ralph Thomas Myers, 2004 **Merrill Chemistry** Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 *Modern Physics* Gary N. Felder, Kenny M. Felder, 2022-09-15 *Modern Physics* intertwines active learning pedagogy with the material typically covered in an introductory survey from the basics of relativity and quantum mechanics through recent developments in particle physics and cosmology The flexible approach taken by the authors allows instructors to easily incorporate as much or as little active learning into their teaching as they choose Chapters are enhanced by Discovery and Active Reading exercises to guide students through key ideas before or during class while ConcepTests help check student understanding and stimulate classroom discussions Each chapter also includes extensive assessment material with a range of basic comprehension questions drill and practice calculations computer based problems and explorations of advanced applications A test bank and interactive animations as well as other support for instructors and students are available online Students are engaged by an accessible and lively writing style thorough explanations Math Interludes which account for varying levels of skill and experience and advanced topics to further pique their interest in physics

Instructor's Resource Manual Silberberg, 1999-07 **Solutions Manual to Accompany General Chemistry, Fourth Edition, and General Chemistry with Qualitative Analysis, Fourth Edition, Whitten, Gailey, Davis** Yi-Noo Tang, Wendy L. Keeney-Kennicutt, 1992 *Reviews in Graph Theory* William G. Brown, 1980 **Topics-aids** Rod O'Connor, 1975 Educational Film/video Locator of the Consortium of University Film Centers and R.R. Bowker Consortium of University Film Centers, 1986 Chemistry Martin Stuart Silberberg, 1996 This general chemistry text

centres on the theme that observable change in chemical systems is the result of molecular change The aims of this edition are to enable students to perceive matter and change at the molecular level and to help build student confidence in their ability to solve chemical problems as they discover the relevance of chemistry to their lives Chemistry George M. Bodner, Harry L. Pardue, 1995 Using an experimental perspective this student friendly textbook teaches chemistry as a process not a product describing research being done in the 90s that relates to material in the book Introduces chemistry in terms of major themes designed to help students build connections between the next series of subjects under consideration and previous chapters Explicit attention is paid to the development of problem solving skills

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